

## Chemistry Worksheets Class 10 on Chapter 1 Chemical Reactions and Equations - Set 3

**Q-1:** Write the equations for the following reactions in the ionic form and balance the ionic equations.

- Magnesium is treated with dilute sulphuric acid
- Stannous chloride solution is added into mercuric chloride solution
- $\text{H}_2\text{S}$  gas is passed into copper(II) sulphate solution.
- Liquid bromine is added to the container of sodium iodide crystals
- Aluminium is treated with dilute hydrochloric acid

**Q-2:** Fill in the blanks

- Aluminium reacts with  $\text{NaOH}$  to form \_\_\_\_\_ and \_\_\_\_\_.
- In the activity series of metal, aluminium is \_\_\_\_\_ active than iron.
- Copper sulphate solution is \_\_\_\_\_ in nature.
- A reaction between an acid and base is called \_\_\_\_\_.
- \_\_\_\_\_ slows down the process of oxidation in packaged food.

**Q-3:** What is the correct representation of the reaction that occurs during the chlor alkali process?

- $2\text{NaCl(l)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(l)} + \text{Cl}_2\text{(g)} + \text{H}_2\text{(g)}$
- $2\text{NaCl(aq)} + 2\text{H}_2\text{O(aq)} \rightarrow 2\text{NaOH(aq)} + \text{Cl}_2\text{(g)} + \text{H}_2\text{(g)}$
- $2\text{NaCl(aq)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{Cl}_2\text{(aq)} + \text{H}_2\text{(aq)}$
- $2\text{NaCl(aq)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{Cl}_2\text{(g)} + \text{H}_2\text{(g)}$

**Q-4:** Which of the following statements is correct?

- Metals act as reducing agent and non metals act as oxidising agent
- Metals act as oxidising agent and non metals acts as reducing agent
- Metals and nonmetals both acts as reducing agents
- Metals and nonmetals both acts as oxidising agents

**Q-5:** What do you expect to happen when the  $\text{AgNO}_3$  solution is added to i)  $\text{NaCl}$  solution ii)  $\text{CCl}_4$ ? Give reasons.

**Q-6:** Give reasons for the following

- Zinc and Aluminium are reactive metals and still they don't corrode easily.
- Bronze metals get coated with a greenish layer if not stored/cleaned properly.

**Q-7:** The following aqueous solutions are given to you:

$\text{ZnSO}_4$ ,  $\text{PbSO}_4$ ,  $\text{Ni}(\text{SO}_4)_2$  and  $\text{Al}_2(\text{SO}_4)_3$

You are also given vessels made of  $\text{Zn}$ ,  $\text{Pb}$ ,  $\text{Al}$  and  $\text{Ni}$ .

# Chapter 1chemical Reactions Practice Problems Answer Key

**JL Elias**



## **Chapter 1 chemical Reactions Practice Problems Answer Key:**

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