

Chemistry Worksheets Class 10 on Chapter 1 Chemical Reactions and Equations- Set 1

Q-1: Write the balanced chemical equations for each of the following reactions.

- Liquid hydrogen peroxide decomposes to form water and oxygen gas.
- Acetylene gas burns in oxygen to form carbon dioxide and water along with evolution of heat.
- When solid mercury(II) oxide is heated, liquid mercury and oxygen gas are produced.
- Marble dissolves in a hydrochloric acid to give calcium chloride, water and carbon dioxide.
- Copper sulphate on treatment with potassium iodide precipitates cuprous iodide, liberates iodine gas and also forms potassium sulphate.

Q-2: Fill in the blanks

- _____ is the magnetic oxide of iron.
- In reactivity series, manganese lies _____ zinc
- To the reaction between AgNO_3 solution and NaCl solution, the names given to this type of reactions are _____ and _____.
- Proteins decompose to form _____ in our body.
- If copper turnings are added to colourless AgNO_3 solution, it turns _____ after sometime.

Q-3: Which of the following displacement reactions will not take place?

- $\text{Zn(s)} + \text{FeSO}_4\text{(aq)} \rightarrow \text{ZnSO}_4\text{(aq)} + \text{Fe(s)}$
- $\text{Fe(s)} + \text{CuSO}_4\text{(aq)} \rightarrow \text{FeSO}_4\text{(aq)} + \text{Cu(s)}$
- $\text{Cu(s)} + 2\text{AgNO}_3\text{(aq)} \rightarrow \text{Cu(NO}_3)_2\text{(aq)} + 2\text{Ag(s)}$
- $\text{Cu(s)} + \text{FeSO}_4\text{(aq)} \rightarrow \text{CuSO}_4\text{(aq)} + \text{Fe(s)}$

Q-4: Which of the following statements is false?

- Oxidation results in the loss of an electropositive element and the gain of an electronegative element.
- Reduction results in the loss of an electronegative element and the gain of an electropositive element.
- A species that gains electrons is referred to as a reducing agent.
- An oxidising agent is a type of specie that absorbs hydrogen atoms.

Q-5: Why are certain chemicals and medications kept in coloured bottles?

Q-6: Sakshi desired that her home be white washed. She went to the market and bought 20 kg of quick lime, which she dissolved in 60 litres of water. When she added lime to water, she noticed that the water began to boil even though it was not heated. Give an explanation for her observation. Name the product and write the corresponding chemical equation and word equation.

Chapter 1self Check Activity Chemistry

**Loretta Jones, Carl Hoeger, Peter
William Atkins, Regina Schoenfield-
Tacher**



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Chemistry: Molecules, Matter, and Change Media Activities Book Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher, 2000-01-15 The Media Activity Book MAB for Jones Atkins Chemistry molecules matter and change contains chapters with lists and descriptions of some of the media available as you study the chapter Each activity begin with a specific textbook reference Then you are given a time estimate of how long it will take to use the media An M media icon in the margin of the textbook means that media exists to support that area of text The media is found in three different places on the website and on two CDs Chemistry 'O' Level Rex M. Heyworth, 2007 Life William K. Purves, 2004 New edition of a text presenting underlying concepts and showing their relevance to medical agricultural and environmental issues Seven chapters discuss the cell information and heredity evolutionary process the evolution of diversity the biology of flowering plants and of animals and ecology and biogeography Topics are linked by themes such as evolution the experimental foundations of knowledge the flow of energy in the living world the application and influence of molecular techniques and human health considerations Includes a CD ROM which covers some of the subject matter and introduces and illustrates 1 700 plus key terms and concepts Annotation copyrighted by Book News Inc Portland OR Essentials of Psychology Robert A. Baron, Michael J. Kalsher, 2002 The paperback essentials of Baron s classic text emphasizes the importance and value of the field of psychology to students showing students how the field can be exported into many areas of their lives The brief edition of this highly praised introductory psychology text continues to present the breadth and scope of psychology in a way that encourages reader involvement and interest Robert Baron is a prize winning textbook author teacher and researcher whose reader friendly writing style draws students into psychology in a way that no other introductory psychology textbook does Resources in Education ,1996 Pharmaceutical Chemistry Chris Rostron, Jill Barber, 2021 Taking medication is a common occurrence for many people whether it is to soothe an aching head regulate blood sugars or to treat life threatening conditions In the UK alone over 900 million prescriptions are dispensed every year Overseeing all of this are pharmacists experts in medicines and their use Pharmaceutical Chemistry provides a wide ranging overview of organic chemistry as applied to the study and practice of pharmacy Drugs are simply chemicals so to fully understand their manufacture formulation and the way they work in our bodies a knowledge of organic compounds and their reactions is essential By reading this book students will begin to understand how a drug molecule is made the process that turns it into a medicine the role the pharmacist has when dispensing that medicine and what happens in the body when it is taken Most importantly the text shows how each of these aspects are integrated helping you to see the bigger picture Pharmaceutical Chemistry is available for students and institutions to purchase in a variety of formats and is supported by online resources The ebook offers a mobile experience and convenient access www.oxfordtextbooks.co.uk/ebooks The online resources include For students Self assessment questions to help the reader to check and reinforce understanding of the material introduced in

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