

(b) At 0°C, the water is at its freezing point, hence both solid and liquid phases are observed.

(c) At 100°C, the water is at its boiling point, hence both liquid and gaseous states of water (water vapour) are observed.

#### Give two reasons to justify –

(a) Water at room temperature is a liquid.

(b) An iron almirah is a solid at room temperature.

Solution:

(a) Water persists as a liquid at room temperature since its melting point is lower than room temperature and its boiling point (100°C) is higher.

Similarly,

(i). A fixed volume is occupied by a fixed mass of water.

(ii). At room temperature, water does not have a fixed shape and flows to fit the container's shape.

As a result, water is a liquid at room temperature.

(b) Because it's melting and boiling points are above room temperature, an iron almirah is a solid at room temperature. In the same way.

(i) An iron almirah is rigid and has a predetermined shape.

(ii) Metals have a relatively high density.

As a result, at room temperature, iron almirah is a solid.

#### 7. Why is ice at 273K more effective in cooling than water at the same temperature?

Solution:

At 273 K, ice will absorb heat energy or latent heat from the medium to overcome fusion and transform into water. As a result, ice has a greater cooling impact than water at the same temperature since water does not absorb the excess heat from the medium.

#### 8. What produces more severe burns, boiling water or steam?

Solution:

Steam produces severe burns. It is because it is an exothermic reaction that releases a high amount of heat which it had consumed during vaporization.

# **Chapter 1states Of Matter Answers**

California, Nathan Newmark

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