

# Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



**Reactant side**

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

**Product side**

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O<sub>2</sub>

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O<sub>2</sub>

molar mass of O<sub>2</sub> = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left( \frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left( \frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left( \frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O<sub>2</sub> to form iron (III) oxide, Fe<sub>2</sub>O<sub>3</sub>. (20pts)

a) Write the balanced chemical equation.



**Reactant side**

4 Fe atoms

3(2) = 6 O atoms

**Product side**

4 Fe atoms

2(3) = 6 O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

**Reactant to each other:**

4 mol Fe: 3 mol O<sub>2</sub>

**Product to each reactant:**

2 mol Fe<sub>2</sub>O<sub>3</sub>: 4 mol Fe

# Chapter 1 stoichiometry Practice Problems

**Ralph Thomas Myers**



## **Chapter 1 stoichiometry Practice Problems:**

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