

## Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
  - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
  - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
  - $\frac{2 \text{ mol Cu}}{1 \text{ mol O}_2} = \frac{1 \text{ mol O}_2}{2 \text{ mol Cu}} = \frac{2 \text{ mol CaO}}{2 \text{ mol Cu}} = \frac{2 \text{ mol Cu}}{2 \text{ mol CaO}} = \frac{1 \text{ mol O}_2}{2 \text{ mol CaO}} = \frac{2 \text{ mol CaO}}{1 \text{ mol O}_2}$  Remember the ratios can be reduced as well!!!
  - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} = \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} = \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol H}_2} = \frac{2 \text{ mol HF}}{2 \text{ mol HF}} = \frac{1 \text{ mol Mg}}{1 \text{ mol Mg}} = \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol H}_2} = \frac{1 \text{ mol H}_2}{1 \text{ mol Mg}}$
- $\frac{185.79 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} = \frac{1 \text{ mol Na}_2\text{CO}_3}{185.79 \text{ g Na}_2\text{CO}_3} = \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} = \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} = \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} = \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} = \frac{100.09 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} = \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3}$
- Molar Mass
  - Definition: the mass of one mole of a substance
  - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen  $\rightarrow$  Water
  - 5.0 mol H<sub>2</sub>
  - 2.5 mol O<sub>2</sub>
- Ethane combustion reaction
  - 15.8 mol O<sub>2</sub>
  - 9.00 mol CO<sub>2</sub> & 13.5 mol H<sub>2</sub>O
- 575 g Na & 886 g Cl<sub>2</sub>
- Iron production reaction
  - 75.1 mol CO
  - 50.1 mol Fe & 75.1 mol CO<sub>2</sub>
- 87,390 g CO or  $8.739 \times 10^4$  g CO & 12,610 g H<sub>2</sub> or  $1.261 \times 10^4$  g H<sub>2</sub>
- Lightning reaction of N<sub>2</sub> & O<sub>2</sub>
  - 1,100 g NO<sub>2</sub> or  $1.10 \times 10^3$  g NO<sub>2</sub>
  720. g NO or  $7.20 \times 10^2$  g NO
- CO<sub>2</sub> removal in spacecraft
  - 42.04 mol NaOH
  - 21.02 mol Na<sub>2</sub>CO<sub>3</sub> & 21.02 mol H<sub>2</sub>O
- Double replacement AgNO<sub>3</sub> & NaBr
  - 463 g NaBr
  - 845 g AgBr
- Soda-acid fire extinguisher
  - 3.058 mol NaHCO<sub>3</sub>
  - 3.058 mol CO<sub>2</sub>, 1.529 mol Na<sub>2</sub>SO<sub>4</sub>, 3.058 mol H<sub>2</sub>O
- Sulfuric acid-sodium hydroxide neutralization reaction
  - H<sub>2</sub>SO<sub>4</sub> + 2 NaOH  $\rightarrow$  Na<sub>2</sub>SO<sub>4</sub> + 2 H<sub>2</sub>O
  - 37 g H<sub>2</sub>SO<sub>4</sub>
  - 53 g Na<sub>2</sub>SO<sub>4</sub> & 14 g H<sub>2</sub>O
- Copper & silver nitrate reaction
  - 0.0104 mol Cu(NO<sub>3</sub>)<sub>2</sub>
  - 0.0104 mol Cu & 0.0209 mol AgNO<sub>3</sub>
- Aspirin production
  - 13.5 kg C<sub>6</sub>H<sub>5</sub>O<sub>4</sub>
  - 7.66 kg C<sub>6</sub>H<sub>5</sub>O<sub>3</sub>
  - 4.29 L HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

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web stimulus for a final writing task in lesson c reading a is a single linear text related to the unit theme each reading passage is recorded on the audio program maps and realistic visuals help to develop learner s visual literacy in preparing to read learners are introduced to key vocabulary items from the reading passage lesson a

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web discuss answers as a class revisit this activity after students have completed the reading answer key preparing to read a 1 aspect 2 spread out 3 income note income can be hourly daily weekly yearly etc 4 productive 5 tend to 6 concentration note another common meaning of concentration is focus 7 suburb b

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web 3 use different parts of speech for instance the paraphrased example above uses the noun birth instead of the adjective born to express the same meaning If the original piece is written in the first person I my team etc you will also need to change the point of view to third person he has a team etc

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