

### Balancing Equations and Simple Stoichiometry-KEY

Balance the following equations:

- 1)  $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
- 2)  $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$
- 3)  $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$
- 4)  $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$
- 5)  $3 \text{ SnO} + 2 \text{ NF}_3 \rightarrow 3 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3$

Using the following equation:



- 6) How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

$$? \text{ g Na}_2\text{SO}_4 = 200 \text{ g NaOH} \rightarrow \text{mol NaOH} \rightarrow \text{mol Na}_2\text{SO}_4 \rightarrow \text{g Na}_2\text{SO}_4$$

$$200 \text{ g NaOH} \times \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} = 5 \text{ mol NaOH} \times \frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} = 2.5 \text{ mol Na}_2\text{SO}_4$$

$$2.5 \text{ mol Na}_2\text{SO}_4 \times \frac{142 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = \boxed{355 \text{ g Na}_2\text{SO}_4}$$

- 7) Using the following equation:



How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

Molar masses:  $\text{LiNO}_3 = 68.95 \text{ g/mol}$  &  $\text{Li}_2\text{SO}_4 = 109.9 \text{ g/mol}$

$$? \text{ g LiNO}_3 = 250 \text{ g Li}_2\text{SO}_4 \rightarrow \text{mol Li}_2\text{SO}_4 \rightarrow \text{mol LiNO}_3 \rightarrow \text{g LiNO}_3$$

$$250 \text{ g Li}_2\text{SO}_4 \times \frac{1 \text{ mol Li}_2\text{SO}_4}{109.9 \text{ g Li}_2\text{SO}_4} = 2.27 \text{ mol Li}_2\text{SO}_4 \times \frac{4 \text{ mol LiNO}_3}{2 \text{ mol Li}_2\text{SO}_4} = 4.54 \text{ mol LiNO}_3$$

$$4.55 \text{ mol LiNO}_3 \times \frac{68.95 \text{ g LiNO}_3}{1 \text{ mol LiNO}_3} = \boxed{314 \text{ g LiNO}_3}$$

# Chapter 1 stoichiometry Worksheet Answer Key

**N Colangelo**



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## **Table of Contents Chapter 1stoichiometry Worksheet Answer Key**

1. Understanding the eBook Chapter 1stoichiometry Worksheet Answer Key
  - The Rise of Digital Reading Chapter 1stoichiometry Worksheet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1stoichiometry Worksheet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 1stoichiometry Worksheet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1stoichiometry Worksheet Answer Key
  - Personalized Recommendations
  - Chapter 1stoichiometry Worksheet Answer Key User Reviews and Ratings
  - Chapter 1stoichiometry Worksheet Answer Key and Bestseller Lists
5. Accessing Chapter 1stoichiometry Worksheet Answer Key Free and Paid eBooks

- Chapter 1stoichiometry Worksheet Answer Key Public Domain eBooks
  - Chapter 1stoichiometry Worksheet Answer Key eBook Subscription Services
  - Chapter 1stoichiometry Worksheet Answer Key Budget-Friendly Options
6. Navigating Chapter 1stoichiometry Worksheet Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Chapter 1stoichiometry Worksheet Answer Key Compatibility with Devices
  - Chapter 1stoichiometry Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chapter 1stoichiometry Worksheet Answer Key
  - Highlighting and Note-Taking Chapter 1stoichiometry Worksheet Answer Key
  - Interactive Elements Chapter 1stoichiometry Worksheet Answer Key
8. Staying Engaged with Chapter 1stoichiometry Worksheet Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 1stoichiometry Worksheet Answer Key
9. Balancing eBooks and Physical Books Chapter 1stoichiometry Worksheet Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 1stoichiometry Worksheet Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 1stoichiometry Worksheet Answer Key
- Setting Reading Goals Chapter 1stoichiometry Worksheet Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 1stoichiometry Worksheet Answer Key
- Fact-Checking eBook Content of Chapter 1stoichiometry Worksheet Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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