

The Mole Study Guide

Mass Conversions: Round your answers to 2 places after the decimal! Make sure to show your steps!

1) There are 3.75 moles of CO_2 .

a) mass

b) mass

c) mass

2) There are 2.5 g of

a) H_2O

b) H_2O

c) H_2O

3) There are 4.5 g of

a) mass

b) mass

4) There are 2.5 g of

a) mass

b) mass

5) There are 2.5 g of

a) mass

b) mass

Name _____ Period _____ Date _____

The Mole Study Guide

Fill in the blank boxes for the image below.



Write the Empirical Formula for each compound. Use subscripts to show the relative ratio of atoms.

Molecular	Empirical
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_4\text{H}_8\text{O}_4$	
$\text{C}_6\text{H}_{10}\text{O}_5$	
$\text{C}_6\text{H}_{12}\text{O}_6$	
$\text{C}_6\text{H}_{12}\text{O}_6$	

Calculate the Molar Mass for the following compounds.

1) H_2O

2) H_2O

3) H_2O

4) H_2O

Calculate the Percent Composition for each element in the following compound.

1) H_2O

2) H_2O

3) H_2O

4) H_2O

5) H_2O

6) H_2O

Determine the Empirical and Molecular Formula for the compound given the information below.

1) A compound has the following percent composition: 48.00% Hg, 27.60% Cl, 24.40% N. The molar mass for Aqueous is 454.0 g/mol.

a) If you have 454 g of Aqueous, how many grams of Hg, Cl, and N do you have?

b) Determine the Empirical Formula for Aqueous.

c) Determine the Molecular Formula for Aqueous.

Chapter 1study Guide Answers The Mole

Pamela Mills,Amina El-Ashmawy



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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