

# Chapter 12

# STUDY GUIDE

## Mixtures and Solutions

### Section 12.1: Homogeneous and Heterogeneous Mixtures

Do your best to answer each question completely and correctly.

The van der Waals equation states that there are three factors that affect the behavior of a gas.

1. The volume of the gas molecules is not negligible compared to the volume of the container.
2. The attractive forces between gas molecules are not negligible.
3. The pressure of the gas is not negligible.
4. The volume of the gas molecules is not negligible compared to the volume of the container.
5. The pressure of the gas is not negligible.
6. The volume of the gas molecules is not negligible compared to the volume of the container.
7. The pressure of the gas is not negligible.
8. The volume of the gas molecules is not negligible compared to the volume of the container.
9. The pressure of the gas is not negligible.
10. The volume of the gas molecules is not negligible compared to the volume of the container.
11. The pressure of the gas is not negligible.
12. The volume of the gas molecules is not negligible compared to the volume of the container.
13. The pressure of the gas is not negligible.
14. The volume of the gas molecules is not negligible compared to the volume of the container.
15. The pressure of the gas is not negligible.
16. The volume of the gas molecules is not negligible compared to the volume of the container.
17. The pressure of the gas is not negligible.
18. The volume of the gas molecules is not negligible compared to the volume of the container.
19. The pressure of the gas is not negligible.
20. The volume of the gas molecules is not negligible compared to the volume of the container.

The van der Waals equation states that there are three factors that affect the behavior of a gas. The van der Waals equation states that there are three factors that affect the behavior of a gas.

Factor	Volume	Pressure	Temperature
1. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
2. The attractive forces between gas molecules are not negligible.	Volume	Pressure	Temperature
3. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
4. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
5. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
6. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
7. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
8. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
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10. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
11. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
12. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
13. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
14. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
15. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
16. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
17. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
18. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature
19. The pressure of the gas is not negligible.	Volume	Pressure	Temperature
20. The volume of the gas molecules is not negligible compared to the volume of the container.	Volume	Pressure	Temperature

# Chapter 1study Guide Mixtures Solutions Answers

**H. Stephen Stoker,Danny V. White**



## **Chapter 1 study Guide Mixtures Solutions Answers:**

*Study Guide/Selected Solutions Manual* Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

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In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Chapter 1study Guide Mixtures Solutions Answers** a literary masterpiece penned by a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of people who partake in its reading experience.

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