

CHAPTER 24 REVIEW

Nuclear Chemistry

SECTION 1

Write a complete answer to the following questions in the space provided.

1. _____ Describe the relationship between the three elementary particles in the atom, as well as their properties/masses?
 - a. Neutrons.
 - b. Electrons.
 - c. Protons.
 - d. They all have similar mass values.
2. _____ What does the atomic number represent regarding to:
 - a. the number nucleons there.
 - b. the number of protons.
 - c. the number of neutrons there.
 - d. the number of electrons there.
3. _____ What property of a nucleus is most affected by the:
 - a. number of neutrons.
 - b. number of protons.
 - c. number of electrons.
 - d. mass of neutrons or protons.
4. _____ How does the mass defect relate to the equation provided?
 - a. mass is lost and energy is added to.
 - b. mass is lost and energy is released.
 - c. mass is gained and energy is added to.
 - d. mass is gained and energy is released.
5. _____ How strong of a gravitational condition does the gravitational binding energy have?
 - a. nuclear binding energy is not involved.
 - b. gravitational binding is not involved.
 - c. the nuclear binding energy is affected as there is a small mass defect.
 - d. nuclear binding energy is not involved.
6. Based on Figure 24.1 of the text, which range of E_{B} values is correct?
 - a. the nuclear binding energy is not involved.
 - b. nuclear binding is not involved.
 - c. the nuclear binding energy is not involved.
 - d. the nuclear binding energy is not involved.
7. The number of neutrons in an atom of hydrogen-2 is _____.
8. How many of the same element have the same _____.

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F Rizvi



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