

Biology Exam #2 (Chapters 2, 3, 4, 6, 7, 8)

Chapter 2:

Atoms are made up of electrons (- outside the nucleus), protons(+inside the nucleus), and neutrons (0 inside the nucleus)

Atomic number= number of protons

Mass number=protons + nucleus

Matter is composed of elements and exists as solid, liquid, gas, or plasma

Element: a substance that cannot be broken down by chemical means to a simpler substance that has different physical properties.

Basic elements to life include C, H, N, O, P, S

Life's diversity is based on carbon, almost all the molecules of a cell are composed of carbon

Organic molecules: carbon-based molecules, molecules have C and H, those with only C and H are called hydrocarbons

Isotopes: atoms of an element with different number of neutrons

Radioactive isotope emits energy as rays and subatomic particles

Uses of radiation include medical imaging, killing bacteria and fungi, killing cancer cells

For neutral atoms, the number of valence electrons is equal to the atom's main group number

ATP: currency of the cell, high-energy molecule, undergoes hydrolysis which releases energy

Polar covalent bonds: sharing between two atoms that is unequal

Oxygen is more electronegative than Hydrogen

Second strongest

Nonpolar covalent bonds: sharing of electrons between two atoms is fairly equal

Strongest

Carbon bonds: four electrons on outer shell

Electronegativity: attraction of an atom for electrons in covalent bond

Hydrogen bonding: attraction between negative oxygen and positive hydrogen

Bonds is weak individually, but strong together[weakest]

Biology 1exam 2 Study Guide

Mark Anestis, Kellie Cox



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described in this book is their emphasis on actively engaging students who seek to make sense of natural objects and events. Many of the strategies we highlight emerge from a constructivist view of learning that has gained widespread acceptance in recent years. In this view, learners make sense of the world by forging connections between new ideas and those that are part of their existing knowledge base. For most students, that knowledge base is riddled with a host of naïve notions, misconceptions, and alternative conceptions they have acquired throughout their lives. To a considerable extent, the job of the teacher is to coax out these ideas to help students understand how their ideas differ from the scientifically accepted view, to assist as students restructure and reconcile their newly acquired knowledge, and to provide opportunities for students to evaluate what they have learned and apply it in novel circumstances. Clearly, this prescription demands far more than most college and university scientists have been prepared for.

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