

## 7. GROUPS 1 & 2

### SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 1 Magnesium oxide, known as magnesia, can be produced by the thermal decomposition of magnesium carbonate.

The equation for the reaction is



- \*(a) Other Group 2 carbonates also undergo thermal decomposition.

State and explain the trend in thermal stability of carbonates down Group 2.

(4)

Thermal stability increases down the group

Because ionic radius increases as you go down the group

This results in less polarisation (distortion) of the carbonate ion

And less weakening of the covalent bonds (C-O) within the carbonate ion

- \*(c) State why carrying out a flame test is not an appropriate way to confirm the presence of magnesium ions in the sample of magnesia.

(1)

Magnesium ions give no flame colour

OR

Energy is emitted with Mg ions is outside of the visible region

(Total for Question 1 = 15 marks)

# Biology B3 Past Papers Edexcel 20 May

**Richard Parsons**



**Biology B3 Past Papers Edexcel 20 May:**

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## **Table of Contents Biology B3 Past Papers Edexcel 20 May**

1. Understanding the eBook Biology B3 Past Papers Edexcel 20 May
  - The Rise of Digital Reading Biology B3 Past Papers Edexcel 20 May
  - Advantages of eBooks Over Traditional Books
2. Identifying Biology B3 Past Papers Edexcel 20 May
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Biology B3 Past Papers Edexcel 20 May
  - User-Friendly Interface
4. Exploring eBook Recommendations from Biology B3 Past Papers Edexcel 20 May
  - Personalized Recommendations
  - Biology B3 Past Papers Edexcel 20 May User Reviews and Ratings
  - Biology B3 Past Papers Edexcel 20 May and Bestseller Lists

5. Accessing Biology B3 Past Papers Edexcel 20 May Free and Paid eBooks
  - Biology B3 Past Papers Edexcel 20 May Public Domain eBooks
  - Biology B3 Past Papers Edexcel 20 May eBook Subscription Services
  - Biology B3 Past Papers Edexcel 20 May Budget-Friendly Options
6. Navigating Biology B3 Past Papers Edexcel 20 May eBook Formats
  - ePub, PDF, MOBI, and More
  - Biology B3 Past Papers Edexcel 20 May Compatibility with Devices
  - Biology B3 Past Papers Edexcel 20 May Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Biology B3 Past Papers Edexcel 20 May
  - Highlighting and Note-Taking Biology B3 Past Papers Edexcel 20 May
  - Interactive Elements Biology B3 Past Papers Edexcel 20 May
8. Staying Engaged with Biology B3 Past Papers Edexcel 20 May
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Biology B3 Past Papers Edexcel 20 May
9. Balancing eBooks and Physical Books Biology B3 Past Papers Edexcel 20 May
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Biology B3 Past Papers Edexcel 20 May
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Biology B3 Past Papers Edexcel 20 May
  - Setting Reading Goals Biology B3 Past Papers Edexcel 20 May
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Biology B3 Past Papers Edexcel 20 May
  - Fact-Checking eBook Content of Biology B3 Past Papers Edexcel 20 May
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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