1. GROUPS 1 & 2

SECTION 8

Answer ALL the questions. Write your answers in the spaces provided.

1 Magnesium oxide, known as magnesia, can be produced by the thermal decomposition of magnesium carbonate.

The equation for the reaction is

$$MgCO_1(s) \rightarrow MgO(s) + CO_2(g)$$

*(a) Other Group 2 carbonates also undergo thermal decomposition.

State and explain the trend in thermal stability of carbonates down Group 2.

(4)

Thermal stability increases down the group

Because lonic radius increases as you go down the group

This results in less polarisation (distortion) of the carbonate ion

And less weakening of the covalent bonds (C-0) within the carbonate ion

*(c) State why carrying out a flame test is **not** an appropriate way to confirm the presence of magnesium ions in the sample of magnesia.

(1)

Magnesium ions give no flame colour

OR

Energy is emitted with Mg ions is outside of the visible region

(Total for Question 1 = 15 marks)

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Arthur James Wells

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web 7 free plans for those individuals who are active in hunting and shooting sports these reloading tables or reloading benches are a must you can save a substantial amount of money by reloading spent casings some of the reloading benches shown require additional equipment for set up