

Rate and Growth Rate $\frac{dY}{dt}$ Population Growth $\frac{dN}{dt} = B - D$ Exponential Growth $\frac{dN}{dt} = r_{\text{max}} N$ Logistic Growth $\frac{dN}{dt} = r_{\text{max}} N \left(\frac{K - N}{K} \right)$ Temperature Coefficient Q_{10}[†] $Q_{10} = \left(\frac{k_2}{k_1} \right)^{\frac{10}{T_2 - T_1}}$ Primary Productivity Calculation $\frac{\text{mg O}_2}{\text{L}} \times \frac{0.698 \text{ mL}}{\text{mg}} = \frac{\text{mL O}_2}{\text{L}}$ $\frac{\text{mL O}_2}{\text{L}} \times \frac{0.536 \text{ mg C fixed}}{\text{mL O}_2} = \frac{\text{mg C fixed}}{\text{L}}$ (at standard temperature and pressure)	dY = amount of change dt = change in time B = birth rate D = death rate N = population size K = carrying capacity r_{max} = maximum per capita growth rate of population T_2 = higher temperature T_1 = lower temperature k_2 = reaction rate at T_2 k_1 = reaction rate at T_1 Q_{10} = the factor by which the reaction rate increases when the temperature is raised by ten degrees	Water Potential (Ψ) $\Psi = \Psi_p + \Psi_s$ Ψ_p = pressure potential Ψ_s = solute potential The water potential will be equal to the solute potential of a solution in an open container because the pressure potential of the solution in an open container is zero. The Solute Potential of a Solution $\Psi_s = -iCRT$ i = ionization constant (this is 1.0 for sucrose because sucrose does not ionize in water) C = molar concentration R = pressure constant ($R = 0.0831$ liter bars/mole K) T = temperature in Kelvin ($^{\circ}\text{C} + 273$)
Surface Area and Volume Volume of a Sphere $V = \frac{4}{3} \pi r^3$ Volume of a Rectangular Solid $V = lwh$ Volume of a Right Cylinder $V = \pi r^2 h$ Surface Area of a Sphere $A = 4\pi r^2$ Surface Area of a Cube $A = 6s^2$ Surface Area of a Rectangular Solid $A = \sum$ surface area of each side	r = radius l = length h = height w = width s = length of one side of a cube A = surface area V = volume Σ = sum of all	Dilution (used to create a dilute solution from a concentrated stock solution) $C_i V_i = C_f V_f$ i = initial (starting) C = concentration of solute f = final (desired) V = volume of solution Gibbs Free Energy $\Delta G = \Delta H - T\Delta S$ ΔG = change in Gibbs free energy ΔS = change in entropy ΔH = change in enthalpy T = absolute temperature (in Kelvin) $\text{pH}^* = -\log_{10} [\text{H}^+]$
* For the purposes of the AP Exam, students will not be required to perform calculations using this equation; however, they must understand the underlying concepts and applications. † For use with labs only (optional).		

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Deborah T. Goldberg



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