



Click on the "Molarity" tab to see how to use the applet. You can also click on the "Help" tab to see the instructions. The applet is designed to help you understand the relationship between the amount of solute and the volume of solution, and how this affects the concentration (molarity) of the solution.



Part 1: Directions: Click "Show values"

Move the sliders up and down to see how they affect molarity.



Molarity is moles per liter, that is, how many moles of solute (orange dots) is dissolved per liter of solution.

Part 2: Determining saturation concentration: how concentrated can you get each solution before the solution is saturated? If you can't saturate the solution, write the largest value you can make. Write the values in the table.

 Cobalt chloride	<input type="text"/>	 Potassium chromate	<input type="text"/>
 Potassium permanganate	<input type="text"/>	 Nickel (II) chloride	<input type="text"/>
 Potassium dichromate	<input type="text"/>	 Copper sulfate	<input type="text"/>
 Gold (III) chloride	<input type="text"/>	 Cobalt (II) nitrate	<input type="text"/>

 Part 2: Determining molarity

Part 2: Determining molarity - drag dots: Using the simulation and the formula for molarity above, calculate the table below using Part 2a.

Moles of Compound (mol)	Liters of Solution (L)	Molarity of Solution (M)	Moles of Compound (mol)	Liters of Solution (L)	Molarity of Solution (M)
.53	.79			.78	.59
.86	.34		.88		1.8
1.0	.20		.95	.84	
.67	.67			.64	.85

Part 2b: Questions:

- Can you dissolve .75 moles of Potassium permanganate ( $\text{KMnO}_4$ ) into 500 mL of water? \_\_\_\_\_, why? / why not? (please show work)
- Can 100 mL of water dissolve 0.25 moles of gold (III) chloride,  $\text{AuCl}_3$ ? \_\_\_\_\_, why? / why not? (please show work)

Doing Further - extrapolating results that you cannot find in the simulation:

- Can 1000 mL of water dissolve 8.0 moles of Copper sulfate ( $\text{CuSO}_4$ )? \_\_\_\_\_, why? / why not? (please show work)
- What is the maximum amount of water to needed to dissolve 0.001 moles of potassium dichromate? (please show work)

# Concentration And Molarity Phet Chemistry Labs Worksheet Answers

**Christian Drosten**



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