

Concept Review

Section: Limiting Reactants and Percentage Yield

Complete each statement below by choosing a term from the following list. Some may be used more than once.

excess limiting limiting actual theoretical
percentage percent theoretical

1. A(n) limiting reactant is involved in and completely used up in a chemical reaction.
2. A(n) excess reactant is involved in reacting but not all is consumed. The amount of product that can be obtained is limited.
3. The amount that is used up in a(n) limiting reactant reaction.
4. The limiting reactant determines how much product can be obtained. There is excess of the reactants that are not involved in the reaction.
5. Limit is a factor in obtaining the theoretical amount.
6. In a(n) limiting reactant reaction, the amount of product that can be obtained is limited by the amount of the limiting reactant. In a(n) excess reactant reaction, the amount of product that can be obtained is limited by the amount of the limiting reactant.
7. The percentage yield is a way to compare actual and theoretical amounts.
8. The percentage yield is calculated as actual yield / theoretical yield times 100.
9. The percentage yield is a way to compare actual and theoretical amounts.
10. The percentage yield is calculated as actual yield / theoretical yield times 100.
11. The percentage yield is a way to compare actual and theoretical amounts.
12. The percentage yield is calculated as actual yield / theoretical yield times 100.

Answer the following items in the space provided.

13. When 100 g of A reacts with 100 g of B, 100 g of C is produced. What is the limiting reactant? What is the theoretical yield of C?



Concept Review Answer Key Limiting Reactants And Percentage Yield

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