

Concentration and Molarity PhET Labs

Name: _____

Procedure: https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all the sliders. Notice what happens to the concentration as solid solute is added and when evaporation occurs.



Concentration

- How does the concentration change as solid solute is added? concentration increases
 - How does the concentration change as additional water is added? concentration decreases
 - How does the concentration change as evaporation occurs? concentration increases
 - How do you know when a solution is saturated? concentration does not change (it also shows on screen)
 - When a solution is saturated, and additional solid solute is added, what happens? solid does not dissolve and settles at bottom
 - Why do you think this is? solvent can't hold any more solute
 - How does adding this additional solute change the concentration of this saturated solution? decreases concentration
 - How does evaporation change the concentration of a saturated solution? no change- solution is already saturated
- Part 2: Concentrated Solutions (same website as part 1 just click on the Solution option)**
- Adding a concentrated solution ... describe a way to determine the concentration of the solution in the spigot. Write your plan here: empty the container and add only one solution from the dropper and then use the sensor to measure concentration



- Using your plan ... how might you get that concentrated solution to become saturated? you can evaporate
- Does your plan work for all the other solutions too? yes Why? / Why Not? evaporation reduces the amount of water (or solute) so concentration increases



Part 3: Molarity https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html

Molarity (M) = $\frac{\text{moles of solute (mol)}}{\text{volume of solution (L)}}$

Molarity is moles per liter, that is, how many moles of solute (entire salt) is dissolved per liter of solution.

First, determine the saturation concentration of each of the solutions, that is, how concentrated can you get each solution before the solution is saturated. If you can't determine the concentration using the simulation "Molarity", try using the simulation "Concentration" (You will use this information again in Part 5, if your instructor requires it)



Molarity

Cobalt (II) nitrate	<u>5.640 M</u>	Potassium chromate	<u>3.350 M</u>
Cobalt chloride	<u>4.330 M</u>	Nickel (II) chloride	<u>5.210 M</u>
Potassium dichromate	<u>0.510 M</u>	Copper sulfate	<u>1.380 M</u>
Gold (III) chloride	<u>2.30 M</u>	Potassium permanganate	<u>0.480 M</u>

Concept Review Concentration And Molarity Answers

Ralph Thomas Myers



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Holt Chemistry Holt Rinehart & Winston, 2003-01-24 Ebook: Chemistry: The Molecular Nature of Matter and Change
Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change **Chemistry: Concepts and Problems** Clifford C. Houk, Richard Post, 1996-03-09 CHEMISTRY SECOND EDITION The fast easy way to master the fundamentals of chemistry Have you ever wondered about the differences between liquids gases and solids Or what actually happens when something burns What exactly is a solution An acid A base This is chemistry the composition and structure of substances composing all matter and how they can be transformed Whether you are studying chemistry for the first time on your own want to refresh your memory for a test or need a little help for a course this concise interactive guide gives you a fresh approach to this fascinating subject This fully up to date edition of Chemistry Concepts and Problems Has been tested rewritten and retested to ensure that you can teach yourself all about chemistry Requires no prerequisites Lets you work at your own pace with a helpful question and answer format Lists objectives for each chapter you can skip ahead or find extra help if you need it Reinforces what you learn with chapter self tests EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS CHANG, 2013-01-07 EBOOK GENERAL CHEMISTRY THE ESSENTIAL CONCEPTS **Holt Chemistry** Ralph Thomas Myers, 2004 **Pharmacy Review**, 1989 **Complete Preparation for the AHPAT, 1999** Williams & Wilkins Review, 1998-04 Here is the most accurate and complete resource designed for students interested in applying for admission to health fields such as medical technology and medical therapy Reviews cover science knowledge verbal and quantitative ability and reading comprehension Students will value the topic outlines for the AHPAT exam the in depth allied health glossary and the full length sample test complete with solutions *Allied Health Professions Admissions Test* Aftab Hassan, Betz Guide, Aftab S Hassan, Ph. D., Williams & Wilkins Lippencott, 1995 This guide is aimed at students interested in allied health careers and covers science knowledge verbal and quantitative ability and reading comprehension based on the methods of active learning The text guides students through a system of self managed self paced study related to the AHPAT Each section offers AHPAT style practice questions and there is also a model examination with answers and explanations The appendix includes a glossary allied health careers *Iterations, II* Russell Batt, John W. Moore, 1987 **Basic Concepts of Chemistry, Study Guide** Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems
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