

# 9

## Thermodynamics

- Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
  - $\Delta S$  is negative and therefore,  $\Delta H$  should be highly positive
  - $\Delta S$  is negative and therefore,  $\Delta H$  should be highly negative
  - $\Delta S$  is positive and therefore,  $\Delta H$  should be negative
  - $\Delta S$  is positive and therefore,  $\Delta H$  should also be highly positive
- For the reaction,  $X_2O_4(l) \longrightarrow 2XO_2(g)$ ,  $\Delta U = 2.1 \text{ kcal}$ ,  $\Delta S = 20 \text{ cal K}^{-1}$  at 300 K. Hence,  $\Delta G$  is [2014]
  - 2.7 kcal
  - 2.7 kcal
  - 9.3 kcal
  - 9.3 kcal
- For a given exothermic reaction,  $K_p$  and  $K_p'$  are the equilibrium constants at temperatures  $T_1$  and  $T_2$ , respectively. Assuming that heat of reaction is constant in temperature range between  $T_1$  and  $T_2$ , it is readily observed that [2014]
  - $K_p > K_p'$
  - $K_p < K_p'$
  - $K_p = K_p'$
  - $K_p = \frac{1}{K_p'}$
- A reaction having equal energies of activation for forward and reverse reactions has [2013]
  - $\Delta S = 0$
  - $\Delta G = 0$
  - $\Delta H = 0$
  - $\Delta H = \Delta G = \Delta S = 0$
- In which of the following reactions, standard reaction entropy changes ( $\Delta S^\circ$ ) is positive and standard Gibbs energy change ( $\Delta G^\circ$ ) decreases sharply with increasing temperature? [2012]
  - $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$
  - $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$
  - $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$
  - $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
- The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
  - 10.52 cal/mol K
  - 21.04 cal/mol K
  - 5.260 cal/mol K
  - 0.526 cal/mol K
- Standard enthalpy of vaporisation  $\Delta_{\text{vap}}H^\circ$  for water at 100°C is 40.66 kJ mol<sup>-1</sup>. The internal energy of vaporisation of water at 100°C (in kJ mol<sup>-1</sup>) is (assume water vapour to behave like an ideal gas). [2012]
  - + 37.56
  - 43.76
  - + 43.76
  - + 40.66
- If the enthalpy change for the transition of liquid water to steam is 30 kJ mol<sup>-1</sup> at 27°C, the entropy change for the process would be [2011]
  - 1.0 J mol<sup>-1</sup>K<sup>-1</sup>
  - 0.1 J mol<sup>-1</sup>K<sup>-1</sup>
  - 100 J mol<sup>-1</sup>K<sup>-1</sup>
  - 10 J mol<sup>-1</sup>K<sup>-1</sup>

# Conceptual Physics Thermodynamics Answers

**Albert A Gayle**



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*Instructor's Manual [to Accompany] Conceptual Physics, Eighth Ed* Paul G. Hewitt, 1998 Conceptual Physics Tenth Edition helps readers connect physics to their everyday experiences and the world around them with additional help on solving more mathematical problems Hewitt's text is famous for engaging readers with analogies and imagery from real world situations that build a strong conceptual understanding of physical principles ranging from classical mechanics to modern physics With this strong foundation readers are better equipped to understand the equations and formulas of physics and motivated to explore the thought provoking exercises and fun projects in each chapter Included in the package is the workbook Mechanics Properties of Matter Heat Sound Electricity and Magnetism Light Atomic and Nuclear Physics Relativity For all readers interested in conceptual physics

**Conceptual Physics**, 1985 Rev ed of Conceptual physics a new introduction to your environment 4th ed 1981

**Theoretical Concepts in Physics** Malcolm S. Longair, 2020-04-16 In this original and integrated approach to theoretical reasoning in physics Malcolm Longair illuminates the subject from the perspective of real physics as practised by research scientists Concentrating on the basic insights attitudes and techniques that are the tools of the modern physicist this approach conveys the intellectual excitement and beauty of the subject Through a series of seven case studies an undergraduate course in classical physics and the discovery of quanta are reviewed from the point of view of how the great discoveries and changes of perspective came about This approach illuminates the intellectual struggles needed to attain understanding of some of the most difficult concepts in physics Longair's highly acclaimed text has been fully revised and includes new studies on the physics of fluids Maxwell's great paper on equations for the electromagnetic field and problems of contemporary cosmology and the very early universe

**Teaching Tips** Marvin Druger, Eleanor Dantzler Siebert, Linda W. Crow, 2004 Like a spirited idea exchange among experienced professors Teaching Tips Innovations in Undergraduate Science Instruction brings you the best thinking about how to engage undergraduate science students Most of the ideas in the book are applicable across the sciences

**The Thermodynamics of Electrical Phenomena in Metals, and A Condensed Collection of Thermodynamic Formulas** Percy Williams Bridgman, 1961

*Waves And Stability In Continuous Media - Proceedings Of The VII Conference* Salvatore Rionero, Tommaso Ruggeri, 1994-10-24 This volume presents an up to date overview of some of the most important topics in waves and stability in continuous media The topics are Discontinuity and Shock Waves Linear and Non Linear Stability in Fluid Dynamics Kinetic Theories and Comparison with Continuum Models Propagation and Non Equilibrium Thermodynamics and Numerical Applications

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now

includes even more technology tools and activities to support differentiated instruction      *Physics Briefs* ,1982

**Principles of Physics** Raymond A. Serway,1997-08      Empirical Philosophy of Science Susann Wagenknecht,Nancy J. Nersessian,Hanne Andersen,2015-06-12 The book examines the emerging approach of using qualitative methods such as interviews and field observations in the philosophy of science Qualitative methods are gaining popularity among philosophers of science as more and more scholars are resorting to empirical work in their study of scientific practices At the same time the results produced through empirical work are quite different from those gained through the kind of introspective conceptual analysis more typical of philosophy This volume explores the benefits and challenges of an empirical philosophy of science and addresses questions such as What do philosophers gain from empirical work How can empirical research help to develop philosophical concepts How do we integrate philosophical frameworks and empirical research What constraints do we accept when choosing an empirical approach What constraints does a pronounced theoretical focus impose on empirical work Nine experts discuss their thoughts and empirical results in the chapters of this book with the aim of providing readers with an answer to these questions      Physics John D. Cutnell,2004      **Conflict Resolution Using the Graph Model:**

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