

Concept Review *continued*

10. What do you notice about all of the ionic charges for Group 1? Group 2? Group 15? Group 16? Group 17?

Group 1 = +1, group 2 = +2, group 15 = -3,
group 16 = -2, group 17 = -1

11. How many valence electrons do atoms in Group 1, Group 2, Group 15, Group 16, and Group 17 have?

1, 2, 3, 4, 5, 6, 7, 8

12. State the octet rule.

The chemical rule of thumb
All atoms want 8 valence electrons

13. Ions of calcium, fluorine, magnesium, and iodine have electron configurations that are similar to which noble gases?

Fluorine - neon, calcium - argon,
magnesium - ~~argon~~^{neon}, iodine - xenon

14. Explain why the properties of ions differ from those of their parent atoms.

have different electron configurations

15. Why do atoms of metals form cations?

metals don't have enough
valence electrons

16. Why do atoms of nonmetals form anions?

non metals are close to an octet

Concept Review Section Phase Equilibrium Answers

Margaret G. Kimble



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