

Name:

SL

[30 Marks]

1. Which conditions must be met for a reaction to take place?

- I. Reactants collide with sufficient energy.
- II. Reactants collide with correct orientation.
- III. Reactants must be in the same state.

↓ not necessary

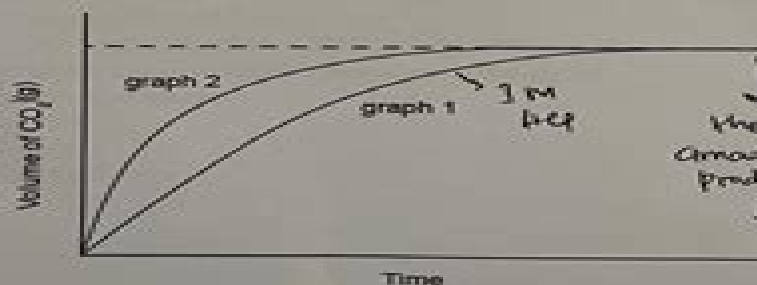
- ☒ A. I and II only
- B. I and III only
- C. II and III only
- D. I, II and III

2. Curves I and II represent samples of the same gas at a constant pressure but at different temperatures. The areas under curves I and II are equal. What does curve II represent?



- A. Curve II is at the lower temperature and there are less molecules in the sample.
- B. Curve II is at the lower temperature and there are the same number of molecules in the samples.
- C. Curve II is at the higher temperature and there are more molecules in the sample.
- ☒ D. Curve II is at the higher temperature and there are the same number of molecules in the samples.

3. Graph 1 shows a plot of volume of $\text{CO}_2(\text{g})$ against time for the reaction of $\text{CaCO}_3(\text{s})$ with $1.00 \text{ mol dm}^{-3} \text{ HCl}(\text{aq})$. The acid is the limiting reagent and entirely covers the lumps of $\text{CaCO}_3(\text{s})$. Which set of conditions is most likely to give the data plotted in graph 2 when the same mass of $\text{CaCO}_3(\text{s})$ is reacted with the same volume of $\text{HCl}(\text{aq})$ at the same temperature?



↓
the same amount of product

↓
the same concentration



↓
limiting reagent

the same mass of CaCO_3

Conceptual Multiple Choice Refraction Questions

Judd E. Hollander



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