

**Chapter 12: Atoms and the Periodic Table**  
**Subatomic Particles**

Three fundamental particles of the atom are the \_\_\_\_\_, \_\_\_\_\_, and \_\_\_\_\_. At the center of each atom lies the atomic \_\_\_\_\_, which consists of \_\_\_\_\_ and \_\_\_\_\_. The atomic number refers to the number of \_\_\_\_\_ in the nucleus. All atoms of the same element have the same number of \_\_\_\_\_, hence, the same atomic number.

Isotopes are atoms that have the same number of \_\_\_\_\_ but a different number of \_\_\_\_\_. An isotope is identified by its atomic mass number, which is the total number of \_\_\_\_\_ and \_\_\_\_\_ in the nucleus. A carbon isotope that has 6 \_\_\_\_\_ and 6 \_\_\_\_\_ is identified as carbon-12, where 12 is the atomic mass number. A carbon isotope having 6 \_\_\_\_\_ and 8 \_\_\_\_\_, on the other hand, is carbon-14.

1. Complete the following table:

Isotope	Number of...		
	Electrons	Protons	Neutrons
Hydrogen-1	1		
Chlorine-36		17	
Nitrogen-14			7
Potassium-40	19		
Arsenic-75		33	
Gold-197			118

2. Which results in a more valuable product — adding or subtracting protons from gold nuclei?

3. Which has more mass, a helium atom or a neon atom?

4. Which has a greater number of atoms, a gram of helium or a gram of neon?



# Conceptual Physical Science Practice Sheet Chapter 12

**G Thomas**



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## **Table of Contents Conceptual Physical Science Practice Sheet Chapter 12**

1. Understanding the eBook Conceptual Physical Science Practice Sheet Chapter 12
  - The Rise of Digital Reading Conceptual Physical Science Practice Sheet Chapter 12
  - Advantages of eBooks Over Traditional Books
2. Identifying Conceptual Physical Science Practice Sheet Chapter 12
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Conceptual Physical Science Practice Sheet Chapter 12
  - User-Friendly Interface
4. Exploring eBook Recommendations from Conceptual Physical Science Practice Sheet Chapter 12
  - Personalized Recommendations
  - Conceptual Physical Science Practice Sheet Chapter 12 User Reviews and Ratings
  - Conceptual Physical Science Practice Sheet Chapter 12 and Bestseller Lists
5. Accessing Conceptual Physical Science Practice Sheet Chapter 12 Free and Paid eBooks
  - Conceptual Physical Science Practice Sheet Chapter 12 Public Domain eBooks
  - Conceptual Physical Science Practice Sheet Chapter 12 eBook Subscription Services
  - Conceptual Physical Science Practice Sheet Chapter 12 Budget-Friendly Options
6. Navigating Conceptual Physical Science Practice Sheet Chapter 12 eBook Formats

- ePub, PDF, MOBI, and More
  - Conceptual Physical Science Practice Sheet Chapter 12 Compatibility with Devices
  - Conceptual Physical Science Practice Sheet Chapter 12 Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Conceptual Physical Science Practice Sheet Chapter 12
    - Highlighting and Note-Taking Conceptual Physical Science Practice Sheet Chapter 12
    - Interactive Elements Conceptual Physical Science Practice Sheet Chapter 12
  8. Staying Engaged with Conceptual Physical Science Practice Sheet Chapter 12
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Conceptual Physical Science Practice Sheet Chapter 12
  9. Balancing eBooks and Physical Books Conceptual Physical Science Practice Sheet Chapter 12
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Conceptual Physical Science Practice Sheet Chapter 12
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Conceptual Physical Science Practice Sheet Chapter 12
    - Setting Reading Goals Conceptual Physical Science Practice Sheet Chapter 12
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Conceptual Physical Science Practice Sheet Chapter 12
    - Fact-Checking eBook Content of Conceptual Physical Science Practice Sheet Chapter 12
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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