

1. Write the equations depicting the behavior of the following acids and bases in water.

- a.  $\text{HBr} \rightarrow \text{H}^+ + \text{Br}^-$
- b.  $\text{HCN} \rightarrow \text{H}^+ + \text{CN}^-$
- c.  $\text{RbOH} \rightarrow \text{Rb}^+ + \text{OH}^-$
- d.  $\text{KOH} \rightarrow \text{K}^+ + \text{OH}^-$

2. Classify each reactant and product as an acid (A), base (B), conjugate acid (CA), or conjugate base (CB).

- a.  $\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$
- b.  $\text{HClO} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{ClO}^-$
- c.  $\text{H}_3\text{PO}_4 + \text{NH}_3 \rightarrow \text{NH}_4^+ + \text{H}_2\text{PO}_4^-$
- d.  $\text{HNO}_3 + \text{HS}^- \rightarrow \text{H}_2\text{S} + \text{NO}_3^-$

a- base, acid, conjugate acid, conjugate base  
b- acid, base, conjugate acid, conjugate base  
c- acid, base, conjugate acid, conjugate base  
d- acid, base, conjugate acid, conjugate base

3. Classify each of the following acids as monoprotic, diprotic, triprotic, or tetraprotic.

- a.  $\text{H}_2\text{SO}_4$  (sulfuric acid) diprotic

- b. Acetic acid, show to the right



monoprotic

- c.  $\text{HCN}$  (cyanic acid) monoprotic

4. Oxaloacetic acid, which is produced in metabolic reactions in the human body has the following structure.



- a. Classify the above structure as a mono-, di-, tri-, or tetraprotic acid. Explain your answer.

Diprotic acid because it can only transfer 2 hydrogen protons

# Ch 9 Practice Problems Chemistry Answers

**JS Bruner**



## Ch 9 Practice Problems Chemistry Answers:

**Survival Guide to General Chemistry** Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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 understand the basic chemical principles related to their career The authors emphasize problem solving and provide a range  
 of practice exercises As in previous editions the text first presents the basic concepts of general chemistry and then moves  
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