

- 2013**
- Q.1 Define Acid, Base and Salt. Also give one example of each.
 Q.2 Calculate the pH and pOH of a 0.1 M Nitric Acid solution.
- 2014**
- Q.1 Define double salt and write the formulae of:
 (i) Potash Alum (ii) Carnallite
 Q.2 Define;
 (i) Standard solution (ii) Acidity of a base (iii) Basicity of acid
 Q.3 Define water of crystallization, write names and chemical formulae of two hydrated salts.
 Q.4 Describe with balanced chemical equations Solvay process for Industrial preparation of Anhydrous Sodium Carbonate.
- 2015**
- Q.1 State Bronsted-Lowry theory of acid and basis with one example of each.
 Q.2 Write chemical formula of the following hydrated salt with their water of crystallization.
 (i) Copper sulphate (ii) Sodium carbonate (iii) Magnesium sulphate
 Q.3 Define neutralization? Write down chemical equation of the reaction of NaOH with HCl and CH_3COOH .
 Q.4 Define pH.
- 2016**
- Q.1 State the Lewis concept about acids and bases and explain it with an example.
 Q.2 Describe the extraction of Sulphur by Frasch process.
- 2017**
- Q.1 Describe the Lewis concept about acids and bases.
 Q.2 Define salts and describe three types of salts with examples.

ACID, BASES AND SALTS

- Q.1 (a) Define the following terms
- | | | |
|--------------------------|-----------------------|--------------------|
| i. Acids | ii. Bases | iii. Strong acids |
| iv. Weak Acids | v. Strong bases | vi. Weak bases |
| vii. Basicity of an acid | viii. Acidity of base | ix. Neutralization |
| x. Titration | xi. pH | xii. pOH |
| xiii. Standard Solution | | |
- Q.1 (b) Write the chemical formulae of:
- | | | |
|------------------|---------------|-------------------|
| (i) Caustic soda | (ii) Soda ash | (iii) Baking soda |
|------------------|---------------|-------------------|

ANS.1.a. Acids: Acids can be defined as

"The compounds which produce hydrogen ions (H^+) in aqueous solution are called acids".

The word acid is derived from a Latin word *acidus*, which means sour. For example HCl , HBr , HI

Bases: Bases can be defined as

"The compounds which produce hydroxide ions (OH^-) in aqueous solution are called bases". For example NaOH , KOH

Strong Acids:

Acids which ionize completely in an aqueous solution are strong acids. For example H_2SO_4 , HCl , HNO_3

Weak Acids:

Acids which do not ionize completely in an aqueous solution are weak acids. For example CH_3COOH , H_2CO_3

Strong Bases:

Bases which ionize completely in aqueous solution are called as strong bases. For example NaOH , KOH

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Table of Contents **Ch Acids Bases Salts Study Guide Answer**

1. Understanding the eBook Ch Acids Bases Salts Study Guide Answer
 - The Rise of Digital Reading Ch Acids Bases Salts Study Guide Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Ch Acids Bases Salts Study Guide Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch Acids Bases Salts Study Guide Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ch Acids Bases Salts Study Guide Answer
 - Personalized Recommendations
 - Ch Acids Bases Salts Study Guide Answer User Reviews and Ratings
 - Ch Acids Bases Salts Study Guide Answer and Bestseller Lists

5. Accessing Ch Acids Bases Salts Study Guide Answer Free and Paid eBooks
 - Ch Acids Bases Salts Study Guide Answer Public Domain eBooks
 - Ch Acids Bases Salts Study Guide Answer eBook Subscription Services
 - Ch Acids Bases Salts Study Guide Answer Budget-Friendly Options
6. Navigating Ch Acids Bases Salts Study Guide Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Ch Acids Bases Salts Study Guide Answer Compatibility with Devices
 - Ch Acids Bases Salts Study Guide Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ch Acids Bases Salts Study Guide Answer
 - Highlighting and Note-Taking Ch Acids Bases Salts Study Guide Answer
 - Interactive Elements Ch Acids Bases Salts Study Guide Answer
8. Staying Engaged with Ch Acids Bases Salts Study Guide Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch Acids Bases Salts Study Guide Answer
9. Balancing eBooks and Physical Books Ch Acids Bases Salts Study Guide Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch Acids Bases Salts Study Guide Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ch Acids Bases Salts Study Guide Answer
 - Setting Reading Goals Ch Acids Bases Salts Study Guide Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ch Acids Bases Salts Study Guide Answer
 - Fact-Checking eBook Content of Ch Acids Bases Salts Study Guide Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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