

Study Tip

Organize New Information Different people learn in different ways. If you learn best visually, draw concept maps, outlines, and diagrams to help you organize the chapter content. If you learn best by listening, take turns reading the chapter aloud with a partner.



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Key Concepts

10.1 The Mole: A Measurement of Matter

- Three methods for measuring the amount of a substance are by count, by mass, and by volume.
- A mole of any substance always contains Avogadro's number of representative particles, or 6.02×10^{23} representative particles.
- The atomic mass of an element expressed in grams is the mass of a mole of the element.
- To calculate the molar mass of a compound, find the number of grams of each element contained in one mole of the compound. Then add the masses of the elements in the compound.

10.2 Mole–Mass and Mole–Volume Relationships

- The molar mass of an element or compound is the conversion factor for converting between the mass and the number of moles of a substance.

- One mole of any gas occupies a volume of 22.4 L at standard temperature and pressure. One mole of any substance contains Avogadro's number of particles, so 22.4 L of any gas at STP contains 6.02×10^{23} representative particles of that gas.

10.3 Percent Composition and Chemical Formulas

- To determine the percent by mass of any element in a given compound, divide the element's mass by the mass of the compound and multiply by 100%.
- An empirical formula of a compound is the simplest whole-number ratio of atoms of the elements in the compound.
- The molecular formula of a compound is either the same as its experimentally determined empirical formula, or it is a simple whole-number multiple of it.

Vocabulary

- | | | |
|----------------------------------|--------------------------------|--|
| • Avogadro's hypothesis (p. 300) | • molar mass (p. 294) | • representative particle (p. 290) |
| • Avogadro's number (p. 290) | • molar volume (p. 300) | • standard temperature and pressure (STP) (p. 300) |
| • empirical formula (p. 300) | • mole (mol) (p. 290) | |
| | • percent composition (p. 300) | |

Key Equations

- moles = representative particles $\times \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ representative particles}}$
- representative particles = moles $\times \frac{6.02 \times 10^{23} \text{ representative particles}}{1 \text{ mole}}$
- mass (grams) = number of moles $\times \frac{\text{mass (grams)}}{1 \text{ mole}}$
- moles = mass (grams) $\times \frac{1 \text{ mole}}{\text{mass (grams)}}$
- $\frac{\text{grams}}{\text{mole}} = \frac{\text{grams}}{1} \times \frac{22.4 \text{ L}}{1 \text{ mole}}$
- volume of gas = moles of gas $\times \frac{22.4 \text{ L}}{1 \text{ mol}}$
- % mass of element = $\frac{\text{mass of element}}{\text{mass of compound}} \times 100\%$

Organizing Information

Organize the major ideas of the chapter.



Concept Map 10: Solve the Concept Map with the help of an interactive tutorial.

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Chapter Resources

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- Core Teaching Resources, Chapter 10, [Practice Problems](#), [Vocabulary Review](#), [Quiz](#), [Chapter Test A](#), [Chapter Test B](#)

Technology

- Computer Test Bank, Chapter 10 Test
- Interactive Textbook with ChemASAP, Chapter 10
- Virtual Chemistry Labs, Lab 11

Chapter 10 Study Guide

David G. Myers, Richard O. Straub



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