



CHAPTER 10

CHEMICAL CALCULATIONS AND CHEMICAL EQUATIONS

Although Chapter 9 was full of questions that began with, “*How much...*,” we are not done with such questions yet. In Chapter 9, our questions focused on chemical formulas. For example, we answered such questions as, “*How much* of the element vanadium can be obtained from 2.3 metric tons of the compound V_2O_5 ?” The chemical formula V_2O_5 told us that there are two moles of vanadium, V, in each mole of V_2O_5 . We used this molar ratio to convert from moles of V_2O_5 to moles of vanadium.

In this chapter, we encounter questions that focus instead on chemical reactions. These questions ask us to convert from amount of one substance in a given chemical reaction to amount of another substance participating in the same reaction. For example, a business manager, budgeting for the production of silicon-based computer chips, wants to know *how much* silicon can be produced from 16 kg of carbon and 32 kg of silica, SiO_2 , in the reaction



A safety engineer in a uranium processing plant, wants to know *how much* water needs to be added to 25 pounds of uranium hexafluoride to maximize the synthesis of UO_2F_2 by the reaction



A chemistry student working in the lab might be asked to calculate *how much* 1-bromo-2-methylpropane, C_4H_9Br , could be made from 6.034 g of 2-methyl-2-propanol, C_4H_9OH , in the reaction



In these calculations, we will be generating conversion factors from the coefficients in the balanced chemical equation.

- 10.1 Equation Stoichiometry
- 10.2 Real-World Applications of Equation Stoichiometry
- 10.3 Molarity and Equation Stoichiometry



How much product can be made from the given reactants?

Chapter 10 Chemical Calculations And Chemical Equations

J. G. R. Briggs



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