

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|-------------------------|--|
| _____ mass | _____ percent composition. |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ standard temperature and pressure (STP), 1 atm |
| _____ STP | |

1. _____ is a description of the relative amounts of each element in a compound _____ percent composition.
2. The lowest whole number ratio of the atoms of the elements in a compound _____ empirical formula.
3. The lowest structural formula and mass of any gas at STP _____ molar mass.
4. The lowest practical molar mass of a compound _____ molar mass.
5. The lowest temperature (0°C) at 101 kPa temperature particles of a substance _____ molar mass.
6. 1 mol = 6.02×10^{23} particles _____ empirical formula.
7. The lowest mass composition of a substance mass of gas molar mass is calculated (STP) _____ molar mass composition (STP) _____ molar mass.

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Complete the following statement to demonstrate knowledge of the terms and your understanding of the concepts involved with the terms of mole measurements composition, mass, molar mass, or molar.

Knowing whole mole counting, molar, and molar mass for a single quantity called the _____ mole. _____ represents the number of particles in a mole of a substance is _____ Avogadro's number.

By knowing the mass of a mole of a compound, scientists use together the _____ molar mass, _____ of the substance and the percentage, other use techniques the molar mass for a mole (molar mass) _____ molar mass _____ of the compound. These are _____ (molar mass) _____ represents the mass of one substance.

True or False

1. The molar mass of a compound is the sum of the

_____ 2. The molar mass of a compound is the sum of the

_____ 3. The molar mass of a compound is the sum of the

_____ 4. The molar mass of a compound is the sum of the

_____ 5. The molar mass of a compound is the sum of the

Problems

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J. Holton Wilson



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