

## Chapter 10: Chemical Quantities

### Section 1 Notes – The Mole: A Measurement of Matter (Read pages 287-296)

#### Mole

- chemist \_\_\_\_\_ for number of \_\_\_\_\_
- equals \_\_\_\_\_ number = \_\_\_\_\_

#### Representative Particles

- atoms – \_\_\_\_\_
- molecules – \_\_\_\_\_
- formula units – \_\_\_\_\_
- ions – \_\_\_\_\_

#### Conversions

- \_\_\_\_\_ mole equals
  - \_\_\_\_\_ atoms, molecules, formula units, ions, things

#### Practice Problems (Do Together)

3. How many moles is  $2.80 \times 10^{24}$  atoms of Si?
4. How many moles if  $2.17 \times 10^{23}$  representative particles of bromine?
5. How many atoms are in 1.14 mol  $\text{SO}_2$ ?
6. How many moles are in  $4.65 \times 10^{24}$  molecules of  $\text{NO}_2$ ?

#### Additional Practice Problems

1. How many moles are equal to  $9.03 \times 10^{24}$  atoms of mercury?
2. How many atoms are in 1.00 mole of sucrose,  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ ?
3. How many atoms of O are in 3.65 moles of  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ ?

# Chapter 10 Chemical Quantities Laboratory Answers

**J Dewey**



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