

Chapter 20 Electrode potentials and electrochemical cells

Practice questions

Question number	Answer	Marks	Guidance
1 (a)	By definition	1	Allow 'set to this value'
1 (b)	1.23V	1	Allow + or -
1 (c)	$\text{Pt} \text{H}_2(\text{g}) \text{OH}^-(\text{aq}), \text{H}_2\text{O}(\text{l}) \text{O}_2(\text{g}) \text{H}_2\text{O}(\text{l}), \text{OH}^-(\text{aq}) \text{Pt}$ Correct but with Pt missing Includes Pt with correct representation	1 1	H_2O not essential, allow reverse order
1 (d)	Uses $\text{O}_2 + 2\text{H}_2\text{O} + 4\text{e}^- \rightarrow 4\text{OH}^-$ And (2x) $2\text{OH}^- + \text{H}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{e}^-$ $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	1 1	
1 (e)	Increases the surface area (so reaction faster)	1	
1 (f)	Overall reaction is the same ($2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$)	1	Or shows e.m.f. is the same
1 (g)	Hydrogen and oxygen supplied continuously OR Can be operated without stopping to recharge	1	Or can be refuelled quickly Allow any one mark
1 (h)	Hydrogen may need to be made using an energy source that is not 'carbon neutral'	1	
2 (a)	Zn	1	This is always the species with the most negative E value (it is oxidised itself).
2 (b) (i)	Fe^{3+}	1	
2 (b) (ii)	Cl_2	1	
2 (c) (i)	The standard electrode potential: 1.25V	1	Remember that: $E_{\text{cell}} = E_{\text{right}} - E_{\text{left}}$
2 (c) (ii)	$\text{Ti}^{3+} + 2\text{Fe}^{2+} \rightarrow 2\text{Fe}^{3+} + \text{Ti}^+$	2	One mark for a balanced equation and one for the correct direction.
3 (a)	$\text{Pt} \text{H}_2 \text{H}^+ \text{Fe}^{2+} \text{Fe}$ Note, allow one mark only for correct symbol in reverse: $\text{Fe} \text{Fe}^{2+} \text{H}^+ \text{H}_2 \text{Pt}$	2	Allow 1 for correct order of symbols but lose 2nd mark for a wrong phase boundary(s) / Pt missing / extra Pt on RHS, additional phase boundary Allow dashed lines for salt bridge Ignore state symbols Ignore 2 if used before H^+
3 (b)	Electron donor	1	Allow (species that) loses electrons Do not allow reference to electron pairs
3 (c)	Cl_2 / chlorine (Species on RHS / electron donor) has most positive / largest E° / has highest potential	1 1	If M1 blank or incorrect cannot score M2 Do not allow reference to e.m.f. or $E(\text{cell})$
3 (d) (i)	Cl / chlorine	1	
3 (d) (ii)	Chlorine +1 to chlorine 0	1	CE if chlorine not identified in 3(d)(i) Allow chlorine +1 to chlorine -1 (in Cl^-) Allow oxidation state decreases by one / two Allow oxidation state changes by -1 OR -2
3 (e)	$4\text{HOCl} + 4\text{H}^+ + 4\text{OH}^- \rightarrow 2\text{Cl}_2 + \text{O}_2 + 6\text{H}_2\text{O}$ OR $4\text{HOCl} \rightarrow 2\text{Cl}_2 + \text{O}_2 + 2\text{H}_2\text{O}$	2	Allow one mark for any incorrect equation that shows $\text{HOCl} \rightarrow \text{Cl}_2 + \text{O}_2$ Allow multiples Ignore state symbols Penalise one mark for uncanceled / uncombined species (eg $\text{H}_2\text{O} + \text{H}_2\text{O}$ instead of $2\text{H}_2\text{O}$)
3 (f)	e.m.f. = $0.40 - (-1.25) = 1.65\text{V}$ / +1.65V	1	Allow -1.65V

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John William Hill, Dorothy M. Feigl



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Organic Chemistry, Student Study Guide and Solutions Manual David R. Klein, 2017-01-04 This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems Readers must learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry Existing textbooks provide extensive coverage of the principles but there is far less emphasis on the skills needed to actually solve problems *Organic Chemistry, 5e Student Study Guide and Solutions Manual* David R. Klein, Laurie S. Starkey, 2025-03-18 Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems With Organic Chemistry Student Study Guide and Solutions Manual 5th Edition students can learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry **Organic Chemistry, 4e Student Solution Manual and Study Guide** David R. Klein, 2021-01-07 Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems With Organic Chemistry Student Solution Manual and Study Guide 4th Edition students can learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry Study Guide and Solutions Manual for Organic Chemistry Digital Update K. Peter C. Vollhardt, Neil E. Schore, 2024-03-25 Written by Organic Chemistry coauthor Neil Schore this invaluable manual includes chapter introductions that highlight new materials chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 *Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition* Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines **The American Catalogue**, 1908 **Organic Chemistry, Study Guide and Solutions Manual** T. W. Graham Solomons, Craig B. Fryhle, 2000 On the cover of this book is a Pacific yew tree found in the ancient forests of the Pacific Northwest The bark of the Pacific yew tree produces Taxol found to be a highly effective drug against ovarian and breast cancer Taxol blocks mitosis during eukaryotic cell division The supply of Taxol from the Pacific yew tree is vanishingly small however A single 100 year old tree provides only about one dose of the

drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

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chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed Paul A. Bartlett, Judith G. Koch, 1992

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