#### Double Replacement Reactions Practice 1.0 Answer New

Objective - Classify chemical reactions and predict the products

- Apply the Law of Conservation of Matter to balance chemical equations

Directions: For each reaction below-

- A. Write the chemical equation, predicting the products as nameded.
- B. Balance the chemical equation.
- E. AgNOs + KsPOs +

EsPOs + 3 AgNOs + 3 ENOs + AgsPOs

2. Calcium acetate + mellion carbonate

EarColloGyle + NasCOs + CaCOs + 2 NaCollyOs

3. Alaninum lodids v mercury(II) nitrate

Z Alba + 3 Hg(NO<sub>2</sub>); - 2 Al(NO<sub>2</sub>); + 3 Hg(s)

4. Potassium chromate + magnesium chiorate

NiCros + MajCloub + MajCros + 2 8/20s

S.: Assemblish carbonate + tron(30) perchlorate

3 (NOS<sub>4</sub>)<sub>2</sub>CO<sub>6</sub> + 2 Pe(CEO<sub>4</sub>)<sub>5</sub> + Pe<sub>2</sub>(CO<sub>5</sub>)<sub>5</sub> + 6 NH<sub>4</sub>CIO<sub>4</sub>

6. Bajoro, - Nir. -

Shalbella - 2 Nory - Shara - 2 North

T. Sodium chlorate + Lead(32) acetate

2 Nath + Philliphiology - 2 Natiothio, + Phillip

B. Ammonium silicute + vanadium(V) sullice

S (NOCASIO) - Valido, - Valido, - Valido, - S (NOCASIO)

# **Double Replacement Practice Reactions Answer Key**

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