# **Practicing Electron Configuration**

1. In the space below, write the unabbreviated electron configuration of the following elements. i. Sodium ii. Iron iii. Bromine iv. Barium v. Radium  2. In the space below, write the abbreviated electron configuration of the following elements. i. Cobalt ii. Silver iii. Tellurium iv. Radium v. Lawrencium  3. Determine what elements are denoted by the following electron configuration. i. 1s² 2s² 2p6 3s² 3p4: ii. 1s² 2s² 2p6 3s² 3p6 4s² 3d° 4p6 5s³: iii. [Kr] 5s² 4d° 5p³: iv. [Xe] 6s² 4f⁴ 5d6: v. [Rn] 7s² 5f*:  4. Answer the following questions i. What is the shape of an s orbital? iii. How many s orbitals can there be in an energy level? iii. How many electrons can occupy an s orbital? iv. What is the shape of a p orbital?	-	
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v. How many p orbitals can there be in an energy level?

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Joacim Rocklöv

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