

Balancing a Chemical Equation – Rules

1. Count your atoms and set up your T-Chart.
2. You can only change coefficients, NEVER change the subscripts.
3. Balance Hydrogen last.
4. Balance Oxygen second-to-last.

Balancing a Chemical Equation – Worked Example



Note – this is a difficult problem. Most require less steps

Step 1: Count your atoms and set up your T-Chart.
Result: Nitrogen and Oxygen are balanced but Hydrogen is not balanced in the products side.



T-Chart Step 1		
Reactants	Atoms	Products
1	Nitrogen	1
2	Oxygen	2 + 1 = 3
3	Hydrogen	2

Step 2: Add more hydrogen to the products by putting a "2" coefficient in front of H_2O .
Result: the increased hydrogen in the products, has more Oxygen and Hydrogen are unbalanced, so adjust down on the reactants side.



T-Chart Step 2		
Reactants	Atoms	Products
2	Nitrogen	1
2	Oxygen	2 + 2 = 4
6	Hydrogen	2 + 2 = 4

Step 3: Add more Oxygen and Oxygen to the reactants by putting a "2" coefficient in front of O_2 and NO .
Result: the increased Oxygen and Nitrogen in the reactants has now everything is balanced, so stop Nitrogen first.



T-Chart Step 3		
Reactants	Atoms	Products
2 + 2 = 4	Nitrogen	2
2 + 2 = 4	Oxygen	2 + 2 = 4
6 + 2 = 8	Hydrogen	2 + 2 = 4

Step 4: Add more Nitrogen to the reactants by putting a "2" coefficient in front of NH_3 .
Result: Nitrogen and Oxygen are balanced but, the increased Hydrogen on the reactants side.



T-Chart Step 4		
Reactants	Atoms	Products
2 + 2 = 4	Nitrogen	2 + 2 = 4
2 + 2 = 4	Oxygen	2 + 2 = 4
6 + 2 = 8	Hydrogen	2 + 2 = 4

Step 5: Add more Nitrogen to the reactants by putting a "2" coefficient in front of NH_3 .
Result: 1 Oxygen to balanced has an increased Oxygen on the reactants side.



T-Chart Step 5		
Reactants	Atoms	Products
2 + 2 = 4	Nitrogen	2 + 2 = 4
2 + 2 = 4	Oxygen	2 + 2 = 4
6 + 2 = 8	Hydrogen	2 + 2 = 4

Balancing Chemical Equations Cheat Sheet

Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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