

Activity 10

Classifying
reactions

On the cards, write

1. Type of chemical reactions and their mechanisms



Introduction: Chemical reactions show how **compounds** and **elements** react with one another. An element is an substance consisting of only one of a kind. Such as the elements of the periodic table (HCl). A compound is a substance made of more than one kind of atoms. Such as water (H₂O) or table salt (NaCl).

Question: How are chemical reactions classified?

1. **Single:** When chemical reactions can be classified as one of four types. Using the chemical equations in the cards on the right, decide the following reactions in the right of reaction.

- | | |
|---|------------------------------|
| <input checked="" type="checkbox"/> A. One reaction is broken down into two or more products. | a. Synthesis |
| <input checked="" type="checkbox"/> A. Aqueous is combined with oxygen to produce carbon dioxide and water. | a. Decomposition |
| <input checked="" type="checkbox"/> A. Two or more separate compounds form one product. | c. Single replacement |
| <input checked="" type="checkbox"/> A. Two compounds react to form two different compounds. | d. Double replacement |
| <input checked="" type="checkbox"/> A. A compound reacts with an element to form a new compound and a different element. | e. Combustion |

2. **Double:** Decide which of the chemical reactions below. Some reactions may already be in balance. In the space to the right, decide the reaction as a synthesis, decomposition, single replacement, or double replacement reaction.

- | | |
|---|-------------------|
| a. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ | Double |
| b. $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}$ | Single |
| c. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ | Synthesis |
| d. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ | Synthesis |
| e. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ | Combustion |

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Harold Querner

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