

Name: _____

Date: _____

Student Exploration: Covalent Bonds

Vocabulary: covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. There are eight markers in a full set, but Flora and Frank each only have seven markers. Flora is missing the red marker, and Frank is missing the blue marker.

What can they do so that each has a full set of markers? **They can share the markers.**

2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set?

They can give the ones they need to each other.

Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the *Covalent Bonds* Gizmo, atoms form bonds in this way.

To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** (▶) to see the electrons orbiting the nucleus of each atom.



1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have? _____7_____
2. Click **Pause** (⏸). Drag an electron from the left atom to the right atom. Click **Play**.
What happens? **They share an electron.**
3. Click **Pause**, drag an electron from the right atom to the left, and then click **Play**.
What happens now? **They share another one.**

Covalent Bond Gizmos Answers

R Sanderson



Covalent Bond Gizmos Answers:

The Covalent Bond Henry Sinclair Pickering, 1977 *Covalent Bond 35 Success Secrets - 35 Most Asked Questions on Covalent Bond - What You Need to Know* Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H bonding Noncovalent bonding Polar Polar covalent bond Polar molecules Covalent bonds Polarity of covalent bonds Polar covalent bond Polarity of molecules Noncovalent bonding Interaction Noncovalent bonding Van der Waals Forces Covalent bonds Subdivision of covalent bonds Covalent Bond Classification Polar covalent bond Hybrids Noncovalent bonding Electrostatic Interactions Polar covalent bond Nonpolar molecules Covalent Bond Classification Other uses Covalent bonds History Noncovalent bonding Dipole Dipole Noncovalent bonding Protein Folding Structure and much more

Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch, 1963 **Chemical Bonds and Bonds Energy** R Sanderson, 2012-12-02 Chemical Bonds and Bonds Energy Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction mechanisms Organized into 12 chapters this edition begins with an overview of the quantitative relationship among three basic properties of an atom namely nonpolar covalent radius electronegativity and homonuclear single covalent bond energy This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electromagnetivity of atoms that form a compound Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds The final chapter deals with the application of the principal ideas and techniques to the oxidation of ethane This book is a valuable resource for organic and inorganic chemists

Atoms & Chemical Bonding Science Learning Guide NewPath Learning, 2014-03-01 The Atoms Atomic Configuration Chemical Bonding Ionic Bonding Ionic Compounds Covalent Bonding Covalent Compounds Naming Compounds and Metallic Bonding

Aligned to Next Generation Science Standards NGSS and other state standards **Chemical Bonds and Bond Energy**
Robert Thomas Sanderson,1971 *Chemistry of the Covalent Bond* Leallyn Burr Clapp,1957 Polar Covalence R
Sanderson,1983-03-28 Polar Covalence provides a detailed account of a successful approach to understanding chemistry
from knowledge of atomic structure and the properties that result from this structure This book discusses the nature of
multiple bonds Organized into 16 chapters this book begins with an overview of the interrelationships of various basic atomic
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reduction of the lone pair bond weakening effect through the formation of multiple bonds The final chapter deals with the
relative roles of principles and practice in the teaching of inorganic and general chemistry This book is a valuable resource
for chemists and students **Acids, Bases, and the Chemistry of the Covalent Bond** Calvin Anthony VanderWerf,1961
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