

ChemThink: Covalent Bonding

Consider bonding between two atoms with _____ electrons _____

When two atoms get close enough, the electrons attract the other atom's

Protons / neutrons / electrons /

Before bonding, the atom's electrons spread out all over their _____

surface _____

the surface of each atom. Once bonded, the

electrons spread out all over their _____ surface _____ the two atoms

Atoms must be able to hold onto their own _____ electrons _____ while

pulling _____

another atom's electrons

Consider bonds from between two _____ atoms _____

When atoms move closer, the potential energy _____ increases / decreases /

At a certain point the potential energy _____ increases _____

it goes to

and shows that atoms closer together than _____ protons _____ in each's

nucleus are _____ repelling _____ each other

The ideal distance between the atoms is known as the _____ bond length _____

Bond's energy is _____ bond energy _____

Bond Type	Draw an Example	# of shared e-	# of e- shared	Strength Potential
Single	<chem>N-H</chem>	<u>1</u>	<u>2</u>	<u>weakest</u>
Double	<chem>O=O</chem>	<u>2</u>	<u>4</u>	_____
Triple	<chem>N#N</chem>	<u>3</u>	<u>6</u>	<u>strongest</u>

Covalent Bonding Answer Key For Chemthink

Leallyn Burr Clapp



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