

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H_2S and H_2O have similar structures and their central atoms belong to the same group. Yet H_2S is a gas at room temperature and H_2O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H_2O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

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Marco Cascella



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Table of Contents Covalent Bonding Worksheet Answers Chapter 8 Pg 19

1. Understanding the eBook Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - The Rise of Digital Reading Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Personalized Recommendations
 - Covalent Bonding Worksheet Answers Chapter 8 Pg 19 User Reviews and Ratings
 - Covalent Bonding Worksheet Answers Chapter 8 Pg 19 and Bestseller Lists
5. Accessing Covalent Bonding Worksheet Answers Chapter 8 Pg 19 Free and Paid eBooks
 - Covalent Bonding Worksheet Answers Chapter 8 Pg 19 Public Domain eBooks
 - Covalent Bonding Worksheet Answers Chapter 8 Pg 19 eBook Subscription Services
 - Covalent Bonding Worksheet Answers Chapter 8 Pg 19 Budget-Friendly Options
6. Navigating Covalent Bonding Worksheet Answers Chapter 8 Pg 19 eBook Formats

- ePub, PDF, MOBI, and More
 - Covalent Bonding Worksheet Answers Chapter 8 Pg 19 Compatibility with Devices
 - Covalent Bonding Worksheet Answers Chapter 8 Pg 19 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Highlighting and Note-Taking Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Interactive Elements Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 8. Staying Engaged with Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 9. Balancing eBooks and Physical Books Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Setting Reading Goals Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Fact-Checking eBook Content of Covalent Bonding Worksheet Answers Chapter 8 Pg 19
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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