

## DRAWING COVALENT BONDS

### Introduction:

A **covalent bond** forms between **nonmetals**. Covalent bonds involve **sharing pairs of electrons** between two atoms so that each atom gains a full outer valence shell. Sharing 1 pair of electrons creates a **single covalent bond**, sharing 2 pairs of electrons creates a **double covalent bond**, and sharing 3 pairs of electrons creates a **triple covalent bond**.



How to Draw Simple Covalent Bond Diagrams	
1.	<b>Draw the electron dot diagrams for each element.</b> <i>HINT: This will help you see the number of unpaired valence electrons for each element.</i>
2.	<b>Count the total number of valence electrons between all atoms.</b> <i>HINT: The total number of valence electrons should be 8 in most cases, except 2 hydrogens. If there are not 8, add more atoms of the element with the least number of valence electrons to start.</i>
3.	<b>Determine which element is the central atom in the molecule, if needed.</b> <i>HINT: The central atom is typically the atom with the largest valence number or the atom that there is only one of in the formula. If there are only 2 atoms in the formula, there is no central atom!</i>
4.	<b>Pair up single electrons from different atoms to form covalent bonds and circle each pair.</b> <i>HINT: Keep doing this until each single electron from all atoms have been paired up. You may need to use double or triple bonds to make this work. This step takes trial and error. If one arrangement is not working, try adding another element or trying a new structure. Practice makes perfect!</i>

### Activity: Drawing Covalent Bond Diagrams:

- For each covalent compound, you have been given the element names and the chemical formula.
- Follow the steps listed above to create a covalent bonding diagram for the compound.
- Remember, this takes trial and error. The first 2 have been done for you as examples.

#	Nonmetal	Nonmetal	Formula	Bonding Diagram
A.	hydrogen $\text{H} \cdot$ <i>(must share 1)</i>	chlorine $\cdot \ddot{\text{Cl}} \cdot$ <i>(must share 1)</i>	$\text{HCl}$ <i>(Formula shows 1 H and 1 Cl)</i>	
B.	hydrogen $\text{H} \cdot$ <i>(must share 1)</i>	oxygen $\cdot \ddot{\text{O}} \cdot$ <i>(must share 2)</i>	$\text{H}_2\text{O}$ <i>(Formula shows 2 H and 1 O)</i>	

# Covalent Bond Notes Ch 8

**Lucas, Corwin**

A red circular graphic with a gradient, appearing as a stylized arrow or a light effect, positioned to the right of the author's name.

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**Solid State Physics** Gerald Burns, 2013-10-22 The objective of Solid State Physics is to introduce college seniors and first year graduate students in physics electrical engineering materials science chemistry and related areas to this diverse and fascinating field I have attempted to present this complex subject matter in a coherent integrated manner emphasizing fundamental scientific ideas to give the student a strong understanding and feel for the physics and the orders of magnitude involved The subject is varied covering many important sophisticated and practical areas which at first may appear unrelated but which are actually built on the same foundation the bonding between atoms the periodic translational symmetry and the resulting electron energy levels The text is comprehensive enough so that the basics of broad areas of present research are covered yet flexible enough so that courses of varying lengths can be satisfied the exercises at the end of each chapter serve to reinforce and extend the text

**Cehmistry Textbook for College and University USA** Ibrahim Sikder, 2023-06-04

**Cehmistry Textbook USA** *An Introduction to Chemistry* Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

**Chemical Misconceptions** Keith Taber, 2002 Part one includes information on some of the key alternative conceptions that have been uncovered by research and general ideas for helping students with the development of scientific conceptions

Student's Guide to Fundamentals of Chemistry, Third Edition, Brescia, Arents, Meislich, Turk Donald Bracken, 1975

Student Solutions Manual to Accompany a Conceptual Introduction to Chemistry Rich Bauer, James Birk, Pamela S. Marks, Pamela Marks, 2006-02 The Student Solutions Manual contains detailed solutions and explanations for all odd numbered problems in the text

**Instructor's Resource Manual** Silberberg, 1999-07 **Student's Guide to Chemistry; a Modern Introduction** David Brooks, 1974

Chemistry Charles E. Mortimer, 1986 **Robinson Chemistry Study Guide** Robinson, 1992 *Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson* Norman E. Griswold, Henry Fuller Holtzclaw, 1988

**The Journal of Materials Education**, 1998

General Chemistry Darrell D. Ebbing, 1986-12 *5 Steps to a 5: AP Chemistry 2024 Elite Student Edition* John T.

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