

COVALENT BONDING

Name _____

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:

*H wants 2 electrons
all other atoms 8*



Note that hydrogen is content with 2, not 8, electrons.

electrons in the middle share

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

wanted available share

<p>1. $\text{H} + \text{H} (\text{H}_2)$</p> <p>W - 4</p> <p>A - 2</p> <p>S - 2</p>	<p><i>wanted 2 + 2</i></p> <p><i>available</i></p> <p>$\text{H} \cdot \cdot \text{H}$</p> <p>$\text{H} \cdot \quad \cdot \text{H}$</p>	<p>$\text{H} - \text{H}$</p> <p><i>single bond = share 2 e⁻</i></p>
<p>2. $\text{F} + \text{F} (\text{F}_2)$</p> <p>W - 8 + 8 = 16</p> <p>A - 7 + 7 = 14</p> <p>S - 2</p>	<p>$\cdot \ddot{\text{F}} \cdot \cdot \ddot{\text{F}} \cdot$</p> <p><i>2 e⁻</i></p>	<p>$\cdot \ddot{\text{F}} - \ddot{\text{F}} \cdot$</p> <p><i>single bond 2 e⁻</i></p>
<p>3. $\text{O} + \text{O} (\text{O}_2)$</p> <p>W - 8 + 8 = 16</p> <p>A - 6 + 6 = 12</p> <p>S - 4</p>	<p>$\cdot \ddot{\text{O}} \cdot \cdot \ddot{\text{O}} \cdot$</p> <p><i>4 e⁻</i></p>	<p>$\cdot \ddot{\text{O}} = \ddot{\text{O}} \cdot$</p> <p><i>4 double bond</i></p>
<p>4. $\text{N} + \text{N} (\text{N}_2)$</p> <p>W - 8 + 8 = 16</p> <p>A - 5 + 5 = 10</p> <p>S - 6</p>	<p>$\cdot \ddot{\text{N}} \cdot \cdot \ddot{\text{N}} \cdot$</p>	<p>$\cdot \ddot{\text{N}} \equiv \ddot{\text{N}} \cdot$</p>
<p>5. $\text{C} + \text{O} (\text{CO}_2)$</p> <p>W - 8 + 8 + 8 = 24</p> <p>A - 4 + 6 + 6 = 16</p> <p>S - 8</p>	<p>$\cdot \ddot{\text{O}} \cdot \cdot \ddot{\text{C}} \cdot \cdot \ddot{\text{O}} \cdot$</p>	<p>$\cdot \ddot{\text{O}} = \text{C} = \ddot{\text{O}} \cdot$</p>
<p>6. $\text{H} + \text{O} (\text{H}_2\text{O})$</p> <p>W - 2 + 2 + 8 = 12</p> <p>A - 1 + 1 + 6 = 8</p> <p>S - 4</p>	<p>$\text{H} \cdot \cdot \ddot{\text{O}} \cdot \cdot \text{H}$</p>	<p>$\text{H} - \ddot{\text{O}} - \text{H}$</p>

Covalent Bonds Practice Answers

David R. Klein, Laurie S. Starkey



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