Student Exploration: Bohr Model of Hydrogen

Vocabulary: absorption spectrum, Bohr model, electron volt, emission spectrum, energy level, ionization energy, laser, orbital, photon

Prior Knowledge Questions (Answer these questions BEFORE using the Gizmo.)

What happens to an electron when it absorbs a tiny packet of light, called a photon?

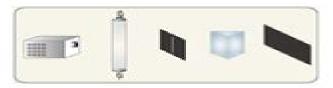
It will move up to a higher energy level.

2. What happens to an electron when it emits a photon?

It will move down to a lower energy level.

Gizmo Warm-up

When white light is passed through hydrogen gas and separated by a prism, some wavelengths of light are absorbed by the hydrogen atoms in the gas. This causes black bars to appear in the absorption spectrum of hydrogen. An emission spectrum is created when hydrogen gas emits light.



The **Bohr model** of the atom was inspired by the spectrum produced by nyurogen gas. The link between light and atomic structure is illustrated by the *Bohr Model of Hydrogen* Gizmo. The Gizmo shows a **laser** pointed at a container of hydrogen gas. After passing through the gas, light from the laser goes through a prism and is detected on a screen.

With the Laser energy set to 7.0 eV, click Play (). Observe the red electron on the

Orbitals pane. Do 7.0-eV photons have any effect on the electron? Photons do not have any effect on the electron as it stays at energy level 1.

- Click Pause (). Set the Laser energy to 12.1 eV and click Play. What happens to the electron now? The electron has moved up to energy level 3.
- 3. Under Go to energy level, select 1. Notice that a photon is emitted by the electron.

What is the energy of the emitted photon? 12.1 eV

Bohr Model Of Hydrogen Gizmo Answer Key

Ricardo Simeoni, Griffith University

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ionization and much more

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Table of Contents Bohr Model Of Hydrogen Gizmo Answer Key

- 1. Understanding the eBook Bohr Model Of Hydrogen Gizmo Answer Key
 - The Rise of Digital Reading Bohr Model Of Hydrogen Gizmo Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Bohr Model Of Hydrogen Gizmo Answer Key
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Bohr Model Of Hydrogen Gizmo Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Bohr Model Of Hydrogen Gizmo Answer Key

- Personalized Recommendations
- Bohr Model Of Hydrogen Gizmo Answer Key User Reviews and Ratings
- Bohr Model Of Hydrogen Gizmo Answer Key and Bestseller Lists
- 5. Accessing Bohr Model Of Hydrogen Gizmo Answer Key Free and Paid eBooks
 - Bohr Model Of Hydrogen Gizmo Answer Key Public Domain eBooks
 - Bohr Model Of Hydrogen Gizmo Answer Key eBook Subscription Services
 - Bohr Model Of Hydrogen Gizmo Answer Key Budget-Friendly Options
- 6. Navigating Bohr Model Of Hydrogen Gizmo Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Bohr Model Of Hydrogen Gizmo Answer Key Compatibility with Devices
 - Bohr Model Of Hydrogen Gizmo Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Bohr Model Of Hydrogen Gizmo Answer Key
 - Highlighting and Note-Taking Bohr Model Of Hydrogen Gizmo Answer Key
 - Interactive Elements Bohr Model Of Hydrogen Gizmo Answer Key
- 8. Staying Engaged with Bohr Model Of Hydrogen Gizmo Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Bohr Model Of Hydrogen Gizmo Answer Key
- 9. Balancing eBooks and Physical Books Bohr Model Of Hydrogen Gizmo Answer Key
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Bohr Model Of Hydrogen Gizmo Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Bohr Model Of Hydrogen Gizmo Answer Key
 - Setting Reading Goals Bohr Model Of Hydrogen Gizmo Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bohr Model Of Hydrogen Gizmo Answer Key

- Fact-Checking eBook Content of Bohr Model Of Hydrogen Gizmo Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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