

# Bond Energy

What makes a reaction endothermic or exothermic?

## Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** ( $\Delta H$ ). In this activity we will look at one way energy changes can be approximated for chemical reactions.

## Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )		Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
  - When bonds are (broken/formed) there is a positive energy change.
  - Breaking bonds is (endothermic/exothermic).
  - When bonds are (broken/formed) there is a negative energy change.
  - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.

- How do the changes in energy for the reverse reactions compare?

*The changes in energy for these reactions are the same, but opposite in sign.*

- Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

*If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.*

# Bond Energy Pogil Extension

**Alice Levine**



## **Bond Energy Pogil Extension:**

**Comprehensive Handbook of Chemical Bond Energies** Yu-Ran Luo, 2007-03-09 Understanding the energy it takes to build or break chemical bonds is essential for scientists and engineers in a wide range of innovative fields including catalysis nanomaterials bioengineering environmental chemistry and space science Reflecting the frequent additions and updates of bond dissociation energy BDE data throughout the literature the Comprehensive Handbook of Chemical Bond Energies compiles the most recent experimental BDE data for more than 19 600 bonds of 102 elements The author organizes the data by bond type functional group bond order bond degree molecular size and structure for ease of use Data can also be located using the Periodic table The book presents data for organic molecules biochemicals and radicals as well as clusters ions hydrogen and surface bonded species van der Waals complexes isotopic species and halogen clusters complexes It also introduces entirely new data for inorganics and organometallics The final chapter summarizes the heats of formation for atoms inorganic organic radicals and monoatomic ions in the gas phase The Comprehensive Handbook of Chemical Bond Energies offers quick access to experimental BDE data in the most inclusive well organized and up to date collection available today CHEMICAL BONDS AND BOND ENERGY. RT. SANDERSON, 1971 Empirical Bond Energies FRANK JR. WILSON, AIR FORCE FLIGHT TEST CENTER EDWARDS AFB CALIF., Air Force Flight Test Center (U.S.), 1961 A consistent set of bond energy values are calculated on the assumption that bond energy nonconstancies are a minimum when determined from compounds in which all the constituent elements exhibit their normal valency Several empirical rules for estimating the heats of vaporization and fusion are included The report contains sufficient data for the estimation of heats of formation of compounds In many cases estimations can be made within a few kilocalories Author **Bond Orders and Energy Components** István Mayer, 2016-10-14 While modern computational methods can provide us with the wave function of a molecule in numerical form most computer programs lack the sophisticated tools needed to extract chemical concepts from these wave functions Saving researchers vast time and potential confusion this volume collects and organizes those validated tools currently scattered throughout the literature and details their application It provides immediate access for those needing to calculate such critical factors as bond order and valence indices and atomic and diatomic contributions to molecular energy Supporting material is available for download from the authors continually updated website **Bond Orders and Energy Components**, 2013 **Municipal Bond Financing for Renewable Energy Systems** Alice Levine, 1981 *Energy Refinements to Glass Structure* Patrick Joseph Hood, 1985 **Bond-based Corrections to Improve Atomization Energy Calculations** Neelabh Maheshwari, 2012 Kit Bond's Position Paper No. 4 Christopher S. Bond, Missouri Transition Government, 1980 Main Group Chemistry, 2013 **Bond Selective Chemistry Beyond the Adiabatic Approximation**, 1993 The adiabatic Born Oppenheimer potential energy surface approximation is not valid for reaction of a wide variety of energetic materials and organic fuels coupling between electronic states of reacting species

plays a key role in determining the selectivity of the chemical reactions induced This research program initially studies this coupling in 1 selective C Br bond fission in 1 3 bromiodopropane 2 C S S H bond fission branching in CH sub 3 SH and 3 competition between bond fission channels and H sub 2 elimination in CH sub 3 NH sub 2 Phase Co-ordinate Bond Graph

Analysis of Multiport Energy Systems ,1974

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**Renewable Energy Systems** Alice Levine,1981 *Phase co-ordinate bond graph analysis of multiport energy systems*

Mohammed Ridzuan Bin Haji Mohammed Salleh,1974

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Improvement and Energy Resources Authority Angela Fick Braly,Beverly A. Marcin,Lewis, Rice & Fingersh,Missouri.

Department of Natural Resources,1994 *S.C.F. Calculations of Bond Properties and Potential Energy Surfaces* ,1971

**Energy and Bonding** Michael Hudson (chimiste),1969

**Bond Selective Chemistry Beyond the Adiabatic**

**Approximation . Technical Progress Report, September 15, 1992--June 14, 1993** ,1993 The adiabatic Born

Oppenheimer potential energy surface approximation is not valid for reaction of a wide variety of energetic materials and organic fuels coupling between electronic states of reacting species plays a key role in determining the selectivity of the chemical reactions induced This research program initially studies this coupling in 1 selective C Br bond fission in 1 3 bromiodopropane 2 C S S H bond fission branching in CH<sub>3</sub>SH and 3 competition between bond fission channels and H<sub>2</sub> elimination in CH<sub>3</sub>NH<sub>2</sub>

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