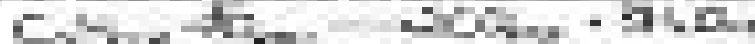


## B = Bonding & Molecular Structure

**Table B.3 • Some Average Single and Multiple Bond Energies (kJ/mol)**

	H	Li	Na	K	Rb	Cs	F	Cl	Br	I	At
H-H	436										
Li-H		404									
Na-H			188								
K-H				163							
Rb-H					147						
Cs-H						147					
F-F							158				
Cl-Cl								243			
Br-Br									193		
I-I										151	
At-At											122
C-C							347				
C-H							413				
C-Cl								339			
C-Br									276		
C-I										238	
C-At											222
C≡C							837				
C≡N							960				
C≡O								1072			
C≡N <sub>2</sub>									891		
C≡O <sub>2</sub>										602	
C≡O <sub>3</sub>											598
C≡O <sub>4</sub>											574
C≡O <sub>5</sub>											550
C≡O <sub>6</sub>											526
C≡O <sub>7</sub>											502
C≡O <sub>8</sub>											478
C≡O <sub>9</sub>											454
C≡O <sub>10</sub>											430
C≡O <sub>11</sub>											406
C≡O <sub>12</sub>											382
C≡O <sub>13</sub>											358
C≡O <sub>14</sub>											334
C≡O <sub>15</sub>											310
C≡O <sub>16</sub>											286
C≡O <sub>17</sub>											262
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C≡O <sub>21</sub>											166
C≡O <sub>22</sub>											142
C≡O <sub>23</sub>											118
C≡O <sub>24</sub>											94
C≡O <sub>25</sub>											70
C≡O <sub>26</sub>											46
C≡O <sub>27</sub>											22
C≡O <sub>28</sub>											-2
C≡O <sub>29</sub>											-26
C≡O <sub>30</sub>											-50
C≡O <sub>31</sub>											-74
C≡O <sub>32</sub>											-98
C≡O <sub>33</sub>											-122
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C≡O <sub>37</sub>											-218
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C≡O <sub>39</sub>											-266
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C≡O <sub>43</sub>											-362
C≡O <sub>44</sub>											-386
C≡O <sub>45</sub>											-410
C≡O <sub>46</sub>											-434
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C≡O <sub>53</sub>											-602
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C≡O <sub>55</sub>											-650
C≡O <sub>56</sub>											-674
C≡O <sub>57</sub>											-698
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C≡O <sub>84</sub>											-1346
C≡O <sub>85</sub>											-1370
C≡O <sub>86</sub>											-1394
C≡O <sub>87</sub>											-1418
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C≡O <sub>93</sub>											-1562
C≡O <sub>94</sub>											-1586
C≡O <sub>95</sub>											-1610
C≡O <sub>96</sub>											-1634
C≡O <sub>97</sub>											-1658
C≡O <sub>98</sub>											-1682
C≡O <sub>99</sub>											-1706
C≡O <sub>100</sub>											-1730

1. Write the balanced equation for the combustion of 1 mole of glucose, C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>, completely in O<sub>2</sub>.



2. Draw the structural formula showing Lewis structures for each of the molecules involved.

B = 1.07, C = 2.55, O = 3.44, H = 2.20, N = 3.04, F = 3.98, Cl = 3.16, Br = 2.96, I = 2.66, At = 2.20



3. Calculate the bond energies of the reactants.

Bond	kJ/mol	kJ/mol
C-H	413	2478
C-O	336	3744
O-O	498	1275

4. Calculate the bond energies of the products.

Bond	kJ/mol	kJ/mol
C=O	799	3794
O-H	463	2728

Total Bond Energy (Reactants)	-4627 kJ	Total Bond Energy (Products)	-5479 kJ
-------------------------------	----------	------------------------------	----------

5. Calculate the Heat Energy released. Bond Energy values.

What is the Heat Energy released (combustion)? -1852 kJ

# Bonding And Molecular Structure Worksheet Answers

## South Pasadena

**Jack Barrett**



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