

Exercises

2. Following the rule you established in the last key question, write correct chemical formula for each of the following compounds

Compound	Formula
(a) Lithium chloride	$\text{Li}^{+1} / \text{Cl}^{-1} \rightarrow \text{LiCl}$
(b) Magnesium iodide	$\text{Mg}^{+2} / \text{I}^{-1} \rightarrow \text{MgI}_2$
(c) Strontium selenide	$\text{Sr}^{+2} / \text{Se}^{-2} \rightarrow \text{SrSe}$
(d) Rubidium fluoride	$\text{Rb}^{+1} / \text{F}^{-1} \rightarrow \text{RbF}$
(e) Lithium oxide	$\text{Li}^{+1} / \text{O}^{-2} \rightarrow \text{Li}_2\text{O}$
(f) Sodium sulfide	$\text{Na}^{+1} / \text{S}^{-2} \rightarrow \text{Na}_2\text{S}$
(g) Potassium chloride	$\text{K}^{+1} / \text{Cl}^{-1} \rightarrow \text{KCl}$
(h) Calcium phosphide	$\text{Ca}^{+2} / \text{P}^{-3} \rightarrow \text{Ca}_3\text{P}_2$
(i) Barium oxide	$\text{Ba}^{+2} / \text{O}^{-2} \rightarrow \text{BaO}$
(j) Aluminum sulfide	$\text{Al}^{+3} / \text{S}^{-2} \rightarrow \text{Al}_2\text{S}_3$

3. Use your answers to the Key Questions and the Exercise Questions to draw a conclusion about the ratio of ions in two compounds if the elements in the compounds are from the same groups (example: aluminum oxide and aluminum sulfide; lithium chloride and potassium chloride).

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