

SECTION - B

Note : There are 12 Short Answer Type Questions from Q. 3 to Q. 14 carrying two marks each. You have to attempt any eight questions. This section includes ten theory questions and two numerical problems. Write appropriate units for each calculated value.

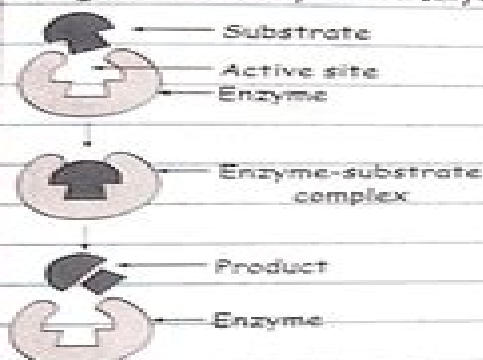
- Q. 3. (a) It oxidises lead sulphide (PbS) to lead sulphate (PbSO_4) changing oxidation state of S from -2 to $+6$.



- (b) Ozone oxidises nitrogen oxide (nitric oxide) to nitrogen dioxide.



- Q. 4. Diagram of enzyme catalysis :



Enzyme catalysis

(2 mark)

- Q. 5. If x_1 and x_2 are the mole fractions of solvent and solute respectively, then

$$x_1 + x_2 = 1$$

By Raoult's law,

$$P = x_1 \times P_0$$

where P^0 is the vapour pressure of a pure solvent and P is the vapour pressure of the solution at given temperature.

$$\therefore \frac{P}{P_0} = x_1$$

$$1 - \frac{P}{P_0} = 1 - x_1 \quad \therefore \frac{P_0 - P}{P_0} = x_2$$

$$\text{If } P_0 - P = \Delta P, \text{ then } \frac{\Delta P}{P_0} = x_2.$$

($\frac{1}{2}$ mark)

($\frac{1}{2}$ mark)

(1 mark)

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Table of Contents Cybered Chemistry Answer Guide

1. Understanding the eBook Cybered Chemistry Answer Guide
 - The Rise of Digital Reading Cybered Chemistry Answer Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Cybered Chemistry Answer Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Cybered Chemistry Answer Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Cybered Chemistry Answer Guide
 - Personalized Recommendations
 - Cybered Chemistry Answer Guide User Reviews and Ratings
 - Cybered Chemistry Answer Guide and Bestseller Lists
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- Cybered Chemistry Answer Guide Public Domain eBooks
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- 6. Navigating Cybered Chemistry Answer Guide eBook Formats
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 - Highlighting and Note-Taking Cybered Chemistry Answer Guide
 - Interactive Elements Cybered Chemistry Answer Guide
- 8. Staying Engaged with Cybered Chemistry Answer Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Cybered Chemistry Answer Guide
- 9. Balancing eBooks and Physical Books Cybered Chemistry Answer Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Cybered Chemistry Answer Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Cybered Chemistry Answer Guide
 - Setting Reading Goals Cybered Chemistry Answer Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Cybered Chemistry Answer Guide
 - Fact-Checking eBook Content of Cybered Chemistry Answer Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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