

Bond Energy

Chem Worksheet 16-2

Name _____

Bond energy is defined as the amount of energy required to break a bond. These values are positive, indicating that bond breaking is endothermic. Bond energies are reported in kilojoules per mole (kJ/mol). The energy for breaking a hydrogen-hydrogen bond is 436 kJ/mol so when a hydrogen-hydrogen bond is formed the process releases 436 kJ/mol.

In a chemical reaction several bonds are broken and formed. For example in the reaction below a hydrogen-hydrogen bond is broken and a fluorine-fluorine bond is broken. Two hydrogen-fluorine bonds are formed. The overall energy change for this process is calculated below.

example



$$\begin{aligned}\Delta H_{\text{rxn}} &= [\text{energy used for breaking bonds}] - [\text{energy formed in making bonds}] \\ &= [436 \text{ kJ/mol} + 155 \text{ kJ/mol}] - [2(567 \text{ kJ/mol})] \\ &= -543 \text{ kJ/mol}\end{aligned}$$



Bond energy: 436 kJ/mol



Bond Energy: 155 kJ/mol



Bond Energy: 567 kJ/mol

Average Bond Energies (kJ/mol)

H-H	436 kJ/mol	C-H	413 kJ/mol	C=C	614 kJ/mol
H-Cl	431 kJ/mol	C-C	348 kJ/mol	C≡C	839 kJ/mol
H-F	567 kJ/mol	C-N	293 kJ/mol	C=O	799 kJ/mol
N-H	391 kJ/mol	C-O	358 kJ/mol	O=O	495 kJ/mol
N-O	201 kJ/mol	C-F	485 kJ/mol	C=O	1072 kJ/mol
O-H	463 kJ/mol	C-Cl	328 kJ/mol	C=N	615 kJ/mol
O-O	146 kJ/mol	C-S	259 kJ/mol	N=N	418 kJ/mol
F-F	155 kJ/mol	Cl-Cl	242 kJ/mol	N≡N	941 kJ/mol
				C≡N	891 kJ/mol

Estimate the enthalpy change (ΔH_{rxn}) of the following reactions using the bond energies above.

- $\text{H}-\text{H} + \text{Cl}-\text{Cl} \rightarrow \text{H}-\text{Cl} + \text{H}-\text{Cl}$
- $\begin{array}{c} \text{H} & & \text{H} \\ & \backslash & / \\ & \text{C}=\text{C} \\ & / & \backslash \\ \text{H} & & \text{H} \end{array} + \text{F}-\text{F} \rightarrow \begin{array}{c} \text{H} & \text{H} \\ | & | \\ \text{H}-\text{C}-\text{C}-\text{H} \\ | & | \\ \text{F} & \text{F} \end{array}$
- $\begin{array}{c} \text{H} \\ | \\ \text{H}-\text{C}-\text{O}-\text{H} \\ | \\ \text{H} \end{array} + \begin{array}{c} \text{H} \\ | \\ \text{H}-\text{C}-\text{O}-\text{H} \\ | \\ \text{H} \end{array} \rightarrow \begin{array}{c} \text{H} & \text{H} \\ | & | \\ \text{H}-\text{C}-\text{O}-\text{C}-\text{H} \\ | & | \\ \text{H} & \text{H} \end{array} + \text{H}-\text{O}-\text{H}$
- $\begin{array}{c} \text{Cl} \\ | \\ \text{H}-\text{C}-\text{Cl} \\ | \\ \text{Cl} \end{array} + \text{O} \rightarrow \begin{array}{c} \text{O} \\ || \\ \text{C}-\text{Cl} \\ | \\ \text{Cl} \end{array} + \text{H}-\text{Cl}$
- $\begin{array}{c} \text{H} \\ | \\ \text{H}-\text{C}-\text{H} \\ | \\ \text{H} \end{array} + \begin{array}{c} \text{O}=\text{O} \\ \text{O}=\text{O} \end{array} \rightarrow \text{O}=\text{C}=\text{O} + \begin{array}{c} \text{H}-\text{O}-\text{H} \\ \text{H}-\text{O}-\text{H} \end{array}$

Draw Lewis structures for reactants and products. Estimate the enthalpy change (ΔH_{rxn}) for the reactions using bond energies.

- $\text{H}_2 (\text{g}) + \text{CO}_2 (\text{g}) \rightarrow \text{H}_2\text{O} (\text{g}) + \text{CO} (\text{g})$
- $2\text{H}_2\text{O}_2 (\text{g}) \rightarrow 2\text{H}_2\text{O} (\text{g}) + \text{O}_2 (\text{g})$
- $\text{CO} (\text{g}) + 2\text{H}_2 (\text{g}) \rightarrow \text{CH}_3\text{OH} (\text{g})$
- $\text{N}_2 (\text{g}) + 3\text{H}_2 (\text{g}) \rightarrow 2\text{NH}_3 (\text{g})$
- $\text{H}_2 (\text{g}) + \text{C}_2\text{H}_4 (\text{g}) \rightarrow \text{C}_2\text{H}_6 (\text{g})$
- $2\text{C}_2\text{H}_6 (\text{g}) + 7\text{O}_2 (\text{g}) \rightarrow 4\text{CO}_2 (\text{g}) + 6\text{H}_2\text{O} (\text{g})$
- $\text{CH}_4 (\text{g}) + 3\text{Cl}_2 (\text{g}) \rightarrow \text{CHCl}_3 (\text{g}) + 3\text{HCl} (\text{g})$
- $\text{HCN} (\text{g}) + 2\text{H}_2 (\text{g}) \rightarrow \text{CH}_3\text{NH}_2 (\text{g})$

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