# **Bond Energy**

What makes a reaction endothermic or exothermic?

## Why?

Chemical reactions can be a lot like playing with Legos\*—you must take apart part of your last creation before you can teplace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new astrongement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the heat of reaction or the change in enthalpy (AH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

# Model 1 - Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mod_)		Reaction	Change in Energy (kJ/mol)
A)	$\mathrm{PCl}_{\mathfrak{g}}(\mathfrak{g}) \ \rightarrow \ \mathbb{P}(\mathfrak{g}) \ + \ 3\mathrm{Cl}(\mathfrak{g})$	+ 966.7 broken	E)	$P(g) + 3Cl(g) \rightarrow PCl_{g}(g)$	- 966.7 formed
B)	$PCl_{g}(g) \rightarrow P(g) + 5Cl(g)$	+ 1.297.9 broken	F)	$P(g) + 5O(g) \rightarrow PO(g)$	- 1,297.9 formed
C)	$PF_3(g) \rightarrow P(g) + 3F(g)$	+ 1,470.4 broken	(2)	$P(g) + 3F(g) \rightarrow PF_3(g)$	- 1,470.4 formed
ED)	$PF_3(g) \rightarrow P(g) + 5F(g)$	+ 2,305.4 broken	H)	P(g) + 5F(g) → PF <sub>3</sub> (g)	- 2,305.4 formed

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed," in the Change in Energy box in the table for each reaction you identify.
- 3. Circle the correct word to complete each sentence below,
  - at. When bonds are (broken) formed) there is a positive energy change.
  - 6. Breaking bunds is (endothermic).
  - c. When honds are (broken/formed) there is a negative energy change.
  - al. Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
  - a. How do the changes in energy for the reverse reactions compare?
    The changes in energy for these reactions are the same, but opposite in sign.
  - Explain your answer to part a considering what you learned from Questions 1—3 about bond breaking and bond formation.

If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.

# **Bond Energy Graph Pogil**

**R Sanderson** 

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<u>Chemistry</u> Ian David Brown,2006 The bond valence model is a recently developed model of the chemical bond in inorganic chemistry that complements the bond model widely used in organic chemistry It is simple quantitative intuitive and predictive no more than a pocket calculator is needed to calculate it This book focuses on the theory that underlies the model and shows how it has been used in physics materials science chemistry mineralogy soil science and molecular biology

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A Pictorial Approach to Molecular Bonding John G. Verkade, 2012-12-06 With the development of **Energy** ,1976 accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non-theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as sociated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics. The major reason for this regrettable situation is the general impracticability of teaching group. theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconcep tion that the ground rules governing bonding in molecules such as NH3 are somehow different from those which apply to aromatic systems such as C H 6 6 Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH for example Introduction to Bond Graphs and Their Applications Jean Ulrich Thoma, 1975 Bond Graph Modelling of Physical Systems Lorcan Stuart Peter Stillwell Smith, 1993 Spatial-energy Principles of the Processes for Complex Structure Formation G. A. Korablev, 2005

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