

## Don't Flip Your Lid: Comparing Intermolecular Forces (from Laying the Foundation)

### Data and Observations

The order of melting (from 1<sup>st</sup> to last): paraffin, iodine, glucose. Sodium chloride does not melt.

### Analysis

*From what you know about intermolecular forces, explain the relative melting points.*

Paraffin is a medium sized nonpolar molecule and iodine is a larger nonpolar molecule. Both paraffin and iodine only contain London forces. However, paraffin has a lower melting point, which makes it melt first. The lower melting point is due to the smaller size of the paraffin molecule. Since iodine is larger, it has more electrons and a larger, more polarizable electron cloud.

Glucose is a medium sized polar molecule that contains London forces. Polar molecules can also form hydrogen bonds. This creates stronger intermolecular forces, producing higher melting points.

Sodium chloride did not melt because it contains ionic bonds, which are very strong. This strength is due to the electrostatic forces within the crystal lattice, which goes in all directions.

### Conclusion

1. Sucrose would have a higher melting point than glucose because it has more electrons. This makes its electron cloud, which is larger, more polarizable. The more polarizable, the more the electron cloud can be distorted, producing a temporary dipole. The larger the dipole created from the polarizability, the larger the intermolecular forces. Larger intermolecular forces mean a higher melting point.
2. Both water and dihydrogen sulfide are bent, polar molecules. However, water has a hydrogen bond between the hydrogen and a very electronegative atom, oxygen. Dihydrogen sulfide do not have these hydrogen bonds, since they only form between hydrogen and electronegative atoms such as oxygen, nitrogen or fluorine.



3. All halogens contain London forces (or dispersion forces) only. London dispersion forces depend on the number of electrons and the size of the electron cloud (greater polarizability). The more electrons a molecule has, the greater the intermolecular forces. The greater the molecular forces, the stronger the attraction between molecules. Iodine has the largest number of electrons, therefore a stronger attraction, creating a solid. Fluorine and chlorine have fewer electrons, creating weaker forces and forming a gas. Bromine is somewhere in between. It's forces are stronger than those in fluorine and chlorine, but not as strong as in iodine.

**Note:** Just because a molecule is bigger (more massive) does not mean it has greater intermolecular forces. This argument confuses gravitational attraction with London forces. Gravitational attraction depends on mass and is very weak and play little to no part in the attraction between molecules. It is the electrostatic forces between molecules due to the distortion of the electron cloud that determines the attraction between molecules.

# Dont Flip Your Lid Comparing Intermolecular Forces

**Geoffrey C. Maitland**



## **Dont Flip Your Lid Comparing Intermolecular Forces:**

**Theory of Intermolecular Forces** H. Margenau, N. R. Kestner, 2013-10-22 Theory of Intermolecular Forces deals with the exposition of the principles and techniques of the theory of intermolecular forces The text focuses on the basic theory and surveys other aspects with particular attention to relevant experiments The initial chapters introduce the reader to the history of intermolecular forces Succeeding chapters present topics on short intermediate and long range atomic interactions properties of Coulomb interactions shape dependent forces between molecules and physical adsorption The book will be of good use to experts and students of quantum mechanics and advanced physical chemistry Intermolecular Forces Pierre L. Huyskens, Werner A.P. Luck, Therese Zeegers-Huyskens, 2012-12-06 The study of intermolecular forces began over one hundred years ago in 1873 with the famous thesis of van der Waals In recent decades knowledge of this field has expanded due to intensive research into both its theoretical and the experimental aspects This is particularly true for the type of very strong cohesive force stressed in 1920 by Latimer and Rodebush the hydrogen bond a phenomenon already outlined in 1912 by Moore and Winemill Hydrogen bonds exert a profound influence on most of the physical and chemical properties of the materials in which they are formed Not only do they govern viscosity and electrical conductivity they also intervene in the chemical reaction path which determines the kinetics of chemical processes The properties of chemical substances depend to a large extent on intermolecular forces In spite of this fundamental fact too little attention is given to these properties both in research and in university teaching For instance in the field of pharmaceutical research about 13000 compounds need to be studied in order to find a single new product that can be successfully marketed The recognition of the need to optimize industrial research efficiency has led to a growing interest in promoting the study of inter molecular forces Rising salary costs in industry have encouraged an interest in theoretical ideas which will lead to tailor made materials **The Theory of Intermolecular Forces** Anthony Stone, 2013-02-01 The theory of intermolecular forces has advanced very greatly in recent years It has become possible to carry out accurate calculations of intermolecular forces for molecules of useful size and to apply the results to important practical applications such as understanding protein structure and function and predicting the structures of molecular crystals The Theory of Intermolecular Forces sets out the mathematical techniques that are needed to describe and calculate intermolecular interactions and to handle the more elaborate mathematical models It describes the methods that are used to calculate them including recent developments in the use of density functional theory and symmetry adapted perturbation theory The use of higher rank multipole moments to describe electrostatic interactions is explained in both Cartesian and spherical tensor formalism and methods that avoid the multipole expansion are also discussed Modern ab initio perturbation theory methods for the calculation of intermolecular interactions are discussed in detail and methods for calculating properties of molecular clusters and condensed matter for comparison with experiment are surveyed **Intermolecular Forces and Their Evaluation by Perturbation Theory** P.

Arrighini,2012-12-06 The aim of these notes is to offer a modern picture of the perturbative approach to the calculation of intermolecular forces The point of view taken is that a perturbative series truncated at a low order can provide a valuable way for valuating interaction energies especially if one limits oneself to the case of intermediate and long range distances between the interacting partners Although the situation corresponding to short distances is essentially left out from our presentation the problems which are within the range of the theory form a vast and important class a large variety of phenomena of matter in fact depends on the existence of interactions among atoms or molecules which over a substantial range of distances should be classified as weak in comparison to the interactions occurring inside atoms or molecules We are aware of the omission of some topics which in principle could have been included in our review For instance a very scarce attention has been paid to the analysis of problems involving interacting partners in degenerate states which is of particular relevance in the case of interactions between excited atoms only a rather quick presentation of the formal apparatus of degenerate perturbation theory is included in Chap III Interactions involving the simultaneous presence of more than two atoms or molecules have not been considered with the consequent non necessity of considering nonadditive effects which characterize the general N body problem

Intermolecular Forces A. Pullman,2013-06-29 Proceedings of the 14th Jerusalem Symposium on Quantum Chemistry and Biochemistry Jerusalem Israel April 13 16 1981

**Theory of Intermolecular Forces** Henry Margenau,N. R. Kestner,1969

*The Theory of Intermolecular Forces* Anthony J. Stone,1996 Describes advances in the theory of intermolecular forces and sets out the mathematical techniques that are needed to handle the more elaborate models that are being used increasingly by both theoreticians and experimentalists Includes a detailed account of the use of higher rank multipole moments to describe electrostatic interactions including treatment of both Cartesian and spherical tensor methods Modern ab initio perturbation theories of intermolecular interactions are also described Annotation copyright by Book News Inc Portland OR

*Intermolecular Forces* Geoffrey C. Maitland,1981

**Cohesion** J. S. Rowlinson,2005-06-30 Why does matter stick together Why do gases condense to liquids and liquids to solids This book provides a detailed historical account of how some of the leading scientists of the past three centuries have tried to answer these questions The topic of cohesion and the study of intermolecular forces has been an important component of physical science research for hundreds of years This book is organised into four broad periods of advances in our understanding The first three are associated with Newton Laplace and van der Waals The final section gives an account of the successful use in the twentieth century of quantum mechanics and statistical mechanics to resolve most of the remaining problems The book will be of primary interest to physical chemists and physicists as well as historians of science interested in the historical origins of our modern day understanding of cohesion

Intermolecular and Surface Forces Jacob N. Israelachvili,2011-07-22 Intermolecular and Surface Forces describes the role of various intermolecular and interparticle forces in determining the properties of simple systems such as gases liquids and solids with a special focus on

more complex colloidal polymeric and biological systems The book provides a thorough foundation in theories and concepts of intermolecular forces allowing researchers and students to recognize which forces are important in any particular system as well as how to control these forces This third edition is expanded into three sections and contains five new chapters over the previous edition Starts from the basics and builds up to more complex systems Covers all aspects of intermolecular and interparticle forces both at the fundamental and applied levels Multidisciplinary approach bringing together and unifying phenomena from different fields This new edition has an expanded Part III and new chapters on non equilibrium dynamic interactions and tribology friction forces      *The Forces Between Molecules* Maurice Rigby,1986 Essential reading for anyone seeking a clear introduction to molecular interactions this book describes the nature of intermolecular forces and their influence on the properties of solids liquids and gases      *Intermolecular Forces* A. Pullman,2014-01-15

**Convergence of Intermolecular Force Series** Glenn M. Roe,1952      *Molecular Interactions* David A. Micha,2020-01-02 A modern comprehensive text and reference describing intermolecular forces this book begins with coverage of the concepts and methods for simpler systems then moves on to more advanced subjects for complex systems emphasizing concepts and methods used in calculations with realistic models and compared with empirical data Contains applications to many physical systems and worked examples Proceeds from introductory material to advanced modern treatments Has relevance for new materials biological phenomena and energy and fuels production      **Intermolecular Forces** General Discussion on Intermolecular Forces,1965      *Intermolecular Forces* Tarō Kihara,1978      **Intermolecular Forces** Taro Kihara,1976      *Intermolecular Forces* Joseph O. Hirschfelder,2009-09-08 The Advances in Chemical Physics series provides the chemical physics and physical chemistry fields with a forum for critical authoritative evaluations of advances in every area of the discipline Filled with cutting edge research reported in a cohesive manner not found elsewhere in the literature each volume of the Advances in Chemical Physics series serves as the perfect supplement to any advanced graduate class devoted to the study of chemical physics      **Intermolecular Forces ,      Intermolecular Forces and Clusters II** D. Wales,2010-11-19

## Whispering the Strategies of Language: An Emotional Journey through **Dont Flip Your Lid Comparing Intermolecular Forces**

In a digitally-driven earth where displays reign great and quick transmission drowns out the subtleties of language, the profound strategies and mental subtleties concealed within words frequently move unheard. However, situated within the pages of **Dont Flip Your Lid Comparing Intermolecular Forces** a captivating fictional prize sporting with organic feelings, lies an extraordinary journey waiting to be undertaken. Published by an experienced wordsmith, that charming opus encourages visitors on an introspective trip, lightly unraveling the veiled truths and profound affect resonating within ab muscles cloth of each and every word. Within the emotional depths of this poignant review, we shall embark upon a honest exploration of the book is core subjects, dissect its charming publishing design, and succumb to the effective resonance it evokes deep within the recesses of readers hearts.

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