



## Nutrition Facts

Serving Size 1 oz. 28g (45g)  
Serving Per Container 32

### Amount Per Serving

Calories 110 Calories from fat 15

% Daily Values\*

Total Fat 2g	3%
Saturated Fat 0g	2%
Trans Fat 0g	0%
Cholesterol 0mg	0%
Sodium 0mg	0%
Total Carbohydrate 30g	12%
Dietary Fiber 3g	

## GOETTA RECIPE

- 1 lb. beef
- 1 lb. pork
- ground together
- 8 cups water
- 2 1/2 cups DORSET'S PINHEAD OATMEAL

- \*1 large onion sliced
- \*1 to 4 bay leaves (optional)
- 3 teaspoons salt a pinch of pepper
- \*VARIATION: Substitute 2 teaspoons dried summer savory in place of onion and bay leaves.

### REGULAR METHOD

1. Put water into pan, when boiling, add salt, pepper, and oatmeal.
2. Cook 2 hours, stirring often, keep lid on while cooking over low heat.
3. Add meat, onion, and bay leaves, mix well.
4. Let cook 1 hour, stirring often.
5. Pour into bread pans. When cool, place in refrigerator. Will keep for days.
6. When ready to use, slice the loaf of goetta and put into a pan in which there is a little hot bacon fat. Fry until well browned. (Very good for breakfast.)

### SLOW COOKER METHOD

- Use same ingredients as in "Regular Method" recipe with ONE EXCEPTION - use only 6 cups water.
1. Put water, salt and pepper into cooker, cover and heat on high for 20 minutes.
  2. Stir in oatmeal, cover and cook on high 1 1/2 hours.
  3. Add meat, onion, and bay leaves - mix well. Cover and cook on low for 3 hours.
  4. Uncover - if not thick enough cook a little longer, stirring often.
  5. Proceed as in direction No.5 and No.6 for "Regular Method".

## OATMEAL CEREAL

### DOUBLE BOILER METHOD

- 3 1/2 cups water
- 1/2 teaspoon salt
- 1 cup DORSET'S PINHEAD OATMEAL

1. Pour water and salt into double boiler and bring to boil.
2. Gradually stir oatmeal into the boiling water.
3. Cover and let cook 1 1/2 to 2 hours, stirring frequently.

# Dorsels Goetta Recipe

**Chao Zhang**



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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal  $\text{KBr}$  that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of  $\text{KBr}$  2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g  $\text{KBr}$

#### **molarity molality and other concentration practice problems** - Nov 23 2021

web calculate the mass percent mole fraction molality and molarity of hexane  $\text{C}_6\text{H}_{14}$  d 0 660 g  $\text{cm}^3$  when it is mixed with heptane  $\text{C}_7\text{H}_{12}$  d 0 684 g  $\text{cm}^3$  in a 35 0 55 0 ml ratio assume that the volumes are additive

#### mole fraction chemteam - Oct 15 2023

web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web a derivatization that achieves this goes as molarity moles of solute volume of solution in liter moles molarity x volume in liter mass moles x molar mass numerical answers 1

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web mole fraction example 1 2m 0 comments mark as completed was this helpful 2 3 problem calculate mole fraction of a 2 4 m aqueous solution of citric acid  $\text{C}_6\text{H}_8\text{O}_7$  a 0 96 b 0 041

*4 7 1 mole practice with guidance exercises - Sep 02 2022*

web solution a  $3.8 \times 10^{-51}$  grams b  $4.11 \times 10^{-45}$  grams c  $7.31 \times 10^{-25}$  grams 4 7 1 mole practice with guidance exercises is shared under a not declared license and was authored remixed and or curated by libretexts these are homework exercises to accompany the textmap created for chemistry a molecular approach by nivaldo tro

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web what is the mole fraction of  $\text{LiCl}$  in the solution 11 8 practice problem acetone  $\text{C}_3\text{H}_6\text{O}$  has a density of  $0.791 \text{ g/ml}$  and xylene  $\text{C}_8\text{H}_{10}$  has a density of  $0.864 \text{ g/ml}$  a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if  $19.0 \text{ ml}$  of acetone was dissolved in  $98.4 \text{ ml}$  of xylene 7 1

chemteam gas law dalton s law problems 1 10 - Feb 24 2022

web since the mole fraction of the  $\text{O}_2$  is  $0.050$  it contributes  $0.050$  or if you will  $5.0$  of the total pressure thus total pressure is  $0.21005042 \text{ atm}$  problem 9 a sample of  $1.43 \text{ g}$  of helium and an unweighed quantity of  $\text{O}_2$  are mixed in a flask at room temperature