

## Covalent Bonding

will need to get more electrons to have a full outer shell. They can do this in 2 ways:

**ionic bonding** - atoms lose or gain electrons, forming positive and negative ions which are then electrostatically attracted to each other (an ionic bond). This occurs between a metal and a non-metal.

**Covalent bonding** - atoms share a pair of electrons (a covalent bond). This occurs between non-metals only.

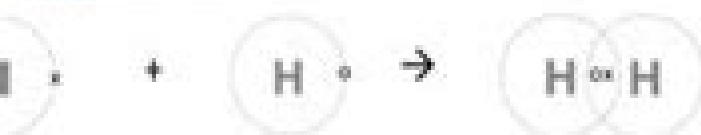
Bonding can only happen between non-metals - it does not involve metals.

atoms 'share' a pair of electrons. They do this by overlapping their outer shells so that the shared electrons are present in shells of both atoms. This shared pair of electrons is a **covalent bond**.



Now that as both atoms now possess the shared pair, both atoms have full outer shells of 8 electrons in their outer shell (in the covalent bond, 8 each on their outer shells).

Hydrogen atoms are involved in covalent bonding, they will only have 2 electrons in their outer shell as the first shell can only hold 2. **Sharing** their outer shell is still full.



**Draw covalent dot cross diagrams:**

How many electrons do the atoms need to gain to have a full outer shell? This is how many bonds (dots) go in the centre.

Fluorine is in group 7, so it has 7 electrons in its outer shell. It needs to gain 1 electron as it will form 1 covalent bond. This is what to draw 1 bond (dot) in the centre.

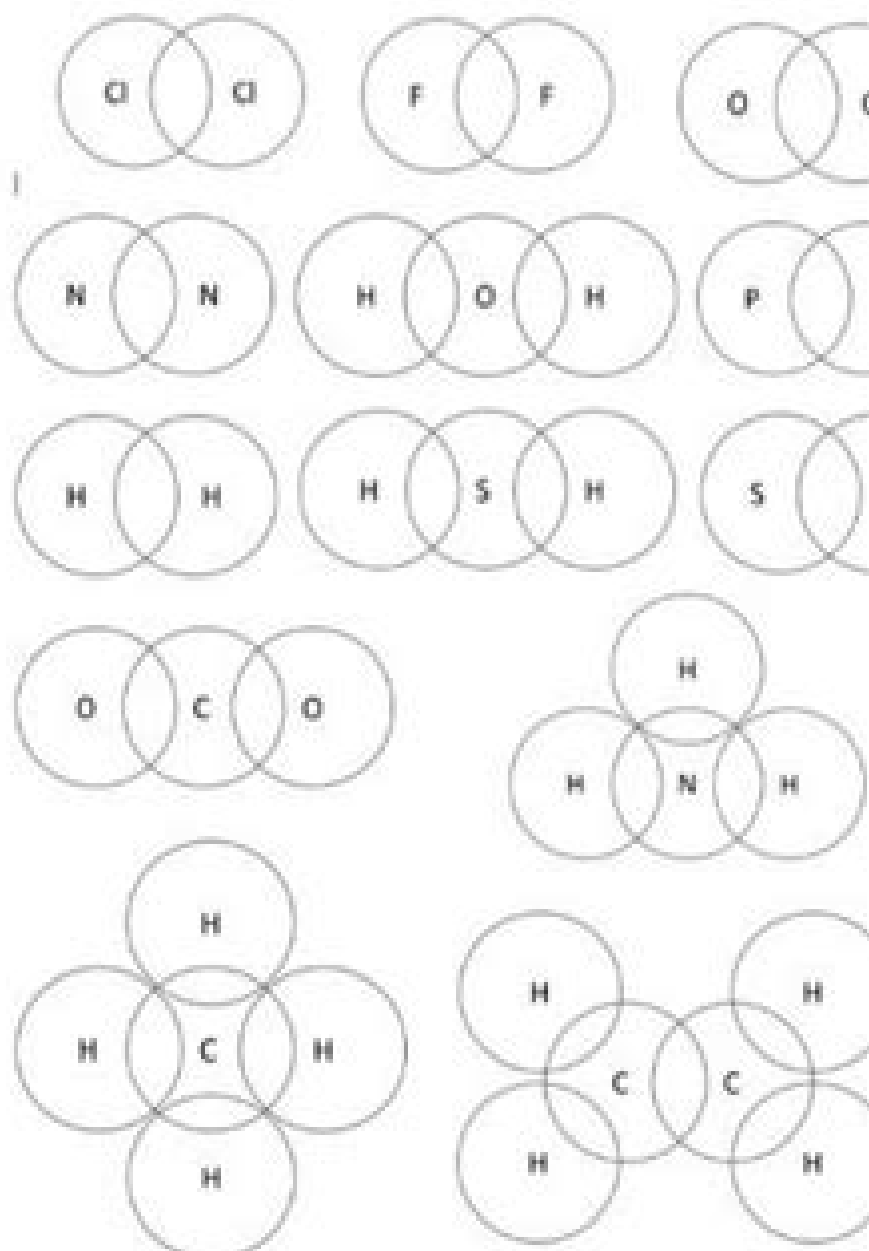


add more electrons to the outer shell of each atom, so that they both have a full outer shell (in this case, 8 electrons). Remember that **both** atoms already 'own' the 2 shared electrons in the centre.

Fluorine already has 2 electrons (from the shared pair), so they each need 6 more. Their outer shells are now full.



## Dot Cross Diagram Practice



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