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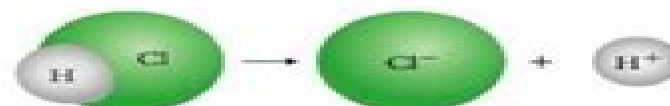
Intro

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Acid

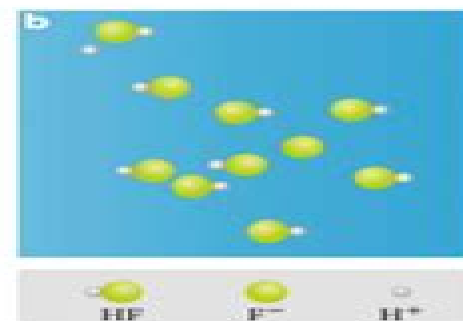
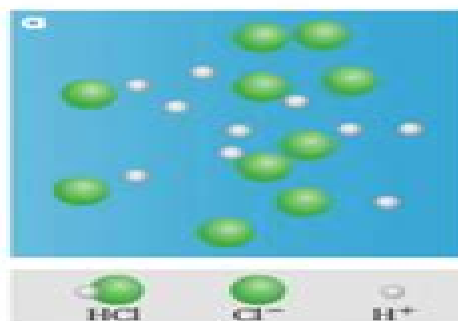
- Produces hydrogen ions
- Or specifically hydronium ions (H_3O^+) in aqueous solution
- Reaction with acid + water \Rightarrow ions are formed \Rightarrow ionisation reaction
- Strong acids \Rightarrow all acid molecules ionise \Rightarrow hydrochloric acid (HCl), Nitric Acid (HNO_3), sulfuric acid (H_2SO_4)

Figure C2.2 ►
Strong acid ionisation:
 $\text{HCl} \rightarrow \text{Cl}^- + \text{H}^+$



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Figure C2.3 ►
Comparison of how
strong acid (HCl) and
weak acid (HF) molecules
ionise in water. a) All of the
HCl molecules are ionised.
b) Only some HF
molecules are ionised.



Base

- Produces hydroxide ion in solution ie. Soluble salts with oxide (O^{2-}) or hydroxide (OH^-) ion
- Soluble base \Rightarrow alkali
- Strong bases always dissociate in water to form separate ions \Rightarrow dissociation reaction

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