

Double Replacement Reactions and Solubility

Post-Lab Questions: Review your Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1) $\text{Na}_2\text{CO}_3(\text{aq}) + \text{Al}(\text{NO}_3)_3(\text{aq}) \rightarrow$	$6\text{NaNO}_3(\text{aq}) + \text{Al}(\text{OH})_3(\text{s})$	None
2) $\text{NaCl}(\text{aq}) + \text{NH}_4\text{NO}_3(\text{aq}) \rightarrow$	$\text{NaNO}_3(\text{aq}) + \text{NH}_4\text{Cl}(\text{aq})$	None
3) $\text{NaI}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$	$\text{NaNO}_3(\text{aq}) + \text{AgI}(\text{s})$	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$
4) $\text{Na}_3\text{PO}_4(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow$	$6\text{NaNO}_3(\text{aq}) + (\text{Cu}_3(\text{PO}_4)_2(\text{s}))$	$3\text{Cu}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Cu}_3(\text{PO}_4)_2(\text{s})$

Use your observations table to answer the following. *ALL ions must include a symbol AND a charge or it is wrong!*

5 a) The symbols and charges of the anions that produced the fewest (<3) precipitates are _____

b) Salts of these anions may be described as SOLUBLE except for those with the cations _____

6a) The symbols and charges of the anion that produced the most (>5) precipitates are _____

b) Salts of these anions may be described as INSOLUBLE except for those with the cations _____

7. Although you did not test the bromide ion (Br^-), would you expect most bromide salts (AlBr_3 , BaBr_2 , CaBr_2 , etc) to be soluble or insoluble? Hint: Locate bromine on the periodic table. _____

8. Explain your answer in #7 _____

9. What cation combined with bromide is an exception to the rule in #7? _____

10-14 Summarize the general solubility rules of ionic compounds in the table below. Two have been done for you.

Ion	Generally soluble or insoluble?	Exceptions
9) carbonates		
10) chlorides, bromides, iodides		
11) hydroxides		
nitrates	Soluble	None
12) phosphates		

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Soil Science Society of America



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General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition. Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works. This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value. E3 Chemistry Guided Study Book - 2018 Home Edition (Answer Key Included) Effiong Eyo, 2017-12-08 Chemistry students and Homeschoolers Go beyond just passing. Enhance your understanding of chemistry and get higher marks on homework quizzes tests and the regents exam with E3 Chemistry Guided Study Book 2018. With E3 Chemistry Guided Study Book students will get clean clear engaging exciting and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry the Physical Setting Easy to read format to help students easily remember key and must know chemistry materials. Several example problems with guided step by step solutions to study and follow. Practice multiple choice and short answer questions along side each concept to immediately test student understanding of the concept. 12 topics of Regents question sets and 2 most recent Regents exams to practice and prep for any Regents Exam. This is the Home Edition of the book. Also available in School Edition ISBN 978 1979088374. The Home Edition contains answer key to all questions in the book. Teachers who want to recommend our Guided Study Book to their students should recommend the Home Edition. Students and parents whose school is not using the Guided Study Book as instructional material as well as homeschoolers should also buy the Home edition. The School Edition does not have the answer key in the book. A separate answer key booklet is provided to teachers with a class order of the book. Whether you are using the school or Home Edition our E3 Chemistry Guided Study Book makes a great supplemental instructional and test prep resource that can be used from the beginning to the end of the school year. PLEASE NOTE Although reading contents in both the school and home editions are identical there are slight differences in question numbers choices and pages between the two editions. Students whose school is using the Guided Study Book as instructional material SHOULD NOT buy the Home Edition. Also available in paperback print. *The Science Teacher's Activity-A-Day, Grades 5-10* Pam Walker, Elaine Wood, 2010-10-05 A hands on and fun filled resource for teaching science to middle and high school students. New in the 5 Minute Fundamentals Series. The Science Teacher's Activity A Day Grades 6-12 includes 180 easy five minute hook or sponge activities to capture learners attention and introduce lessons. Divided into three units Physical Science Life Science and Earth and Space Science the activities cover topics based on the National Science Education Standards. All the book's activities can be done with materials that are inexpensive and easy to find. Includes quick and fun sponge activities that are designed to engage students. All the activities take about 5 minutes to complete. The Science Teacher's Activity a Day is an ideal resource for middle and high school science teachers. **General, Organic, and Biochemistry** Katherine J. Denniston, Joseph J. Topping, Robert L. Caret, 2001 **AP Chemistry Premium,**

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Chemistry Charles H. Corwin, 1994 The book focuses on the concepts of chemistry and the applications that maintain and generate motivation for the subject of chemistry

AP Chemistry Premium, 2024: 6 Practice Tests + Comprehensive Review + Online Practice Neil D. Jespersen, Pamela Kerrigan, 2023-07-04 A guide to taking the Advanced Placement exam in chemistry featuring a review of major chemistry concepts practice and diagnostic tests test taking strategies an overview of the test and practice problems

How to Prepare for the AP Chemistry Neil D. Jespersen, Neil Jespersen, PH. D., 1999-02 This manual offers two diagnostic exams to help students pinpoint their strengths and weaknesses plus three full length practice exams closely modeled on actual AP chemistry exams

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