

ATOMIC STRUCTURE

a guided reading

Vocabulary

Element:

Proton:

Neutron:

Atomic Mass:

Electron:

Atomic Number:

Atomic Structure

There are over 100 different types of atoms, called **elements**. Each atom is made up of three smaller particles called **protons**, **neutrons**, and **electrons**.



Protons and **neutrons** are found in the center of the atom in a space called the **nucleus**. Protons are positively charged, and neutrons do not have a charge- they are neutral. Because of this, the nucleus has an overall positive charge. Ernest Rutherford discovered this in 1899 when he observed that other positively charged particles were repelled by the nucleus. Because protons and neutrons are much heavier than electrons, these particles make up the **atomic mass** of the atom.

Electrons are found outside of the nucleus moving around in orbital "clouds". Scientists used to think that they orbited the nucleus in a ring like pattern, but we now know electrons do not orbit. The electrons are so small and distant from the nucleus that most of the atom is just empty space!

Electrons have a negative charge. An atom has an overall charge of zero when the number of negative electrons and positive protons are equal. However, many chemical reactions involve changing the number of electrons of an atom. When electrons are gained or lost, the atom is called an **ion**. The number of neutrons can be different too between atoms of the same element. Atoms of the same element with different numbers of neutrons are called **isotopes**.

The number of protons never changes. The amount of protons in the nucleus, called the **atomic number**, is what gives an atom its identity as a particular element. Each element on the Periodic Table has a different number of protons, and it is the basis for how the elements are arranged.

Reading Comprehension

1. What are the three particles within an atom?
2. Which two particles are found in the nucleus?
3. Where are the electrons found in an atom?
4. Which particle gives an atom its identity as a particular element?
5. How are elements arranged on the Periodic Table?

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Judd E. Hollander



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