Chapter 9 Practice Problems: Chemical Names and Formulas Section 9.1: Naming Ions

- 1. What is the charge on the ion typically formed by each element?
 - e. nickel, two electrons lost a. oxygen c. sodium
 - b. iodine d. aluminum f. magnesium
- 2. How many electrons does the neutral atom gain or lose when each ion forms?
 - a. Cr c. Li* c. CIT
 - d. Ca2+ b. p3-£ O2-
- Name each ion. Identify each as a cation or an anion.
 - n Sn2* c. Br e 147
 - b. Co3+ d. K* f. Man 2"
- 4. Write the formula (including the charge) for each ion. Use table 9.3 as if necessary.
 - a. carbonate ion c. sulfate ion e. chromate ion
 - b. nitrite ion d. hydroxide ion f. ammonium ion
- 5. Name each of the following ions. Identify each as a cation or an anion.
 - c. Ca2+ a. CNc. POv3-
 - f. SO₂2d. CIT Б. НСО»

9.2 Naming and Writing Formulas for Ionic Compounds:

- Write the formulas for these binary ionic compounds.
 - a. magnesium oxide c. potassium iodide
 - d. aluminum chloride f. ferrie bromide b. tin(II) fluoride

e. sodium sulfide

- 2. Write the formulas for the compounds formed from these pairs of ions.
 - a. Ba^{2*} and Cl⁺ c. Ca2+ and S2-
 - e. Al3+ and O2-
 - b. Ag" and I" d. K' and Br f. Fe2" and O2-
- Name the following binary ionic compounds.
 - a. MnOs c. CaCl₂ e. NiCh g. FenNa
 - b. LisN d. SrBry f. KaS b. SnCl_i
- Write formulas for the following ionic compounds.
 - a. sodium phosphate c. sodium hydroxide e. ammonium chloride
 - b. magnesium sulfate d. potassium evanide f. potassium dichromate
- Write formulas for the following ionic compounds.
 - a. NH₄* and SO₄2c. barium ion and hydroxide ion
 - b. K" and NO: d. lithium ion and carbonate ion

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Jay Ellis Stannard

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